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Synthesis of Embryonic Zeolites with Controlled Physicochemical Properties

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ABSTRACT: Colloidal zeolite precursors, with a sharp particle size distribution (ca. 3-5 nm) and identical chemical composition, the so-called embryonic zeolites (EZs), are prepared by a strict control of aluminosilicate precursors polymerization in zeolite-yielding systems. The organic structure directing agent (OSDA) acts as a sacrificial template and is eliminated by high temperature combustion after the synthesis. Physicochemical properties of the EZs, such as pore size and volume and specific surface area are determined by the size of the OSDA and the synthesis conditions employed; i.e. the larger the OSDA, the higher the microporous volume and the specific surface area of the derived EZs. The EZs belong to the family of extra-large microporous (1-2.5 nm) materials. Upon calcination EZs retain their structure/porosity, and most of their aluminum remains in tetrahedral coordination to provide Brønsted acidity. Pyridine adsorption shows a lower acidic strength for embryonic zeolites with respect to their crystalline counterparts (zeolites). An appropriate combination of extra-large micropores (1-2.5 nm) and Brønsted acid sites (~25 μmol.g⁻¹) leads to improved catalytic performances in the dealkylation of TiPBz, a proxy for bulky molecules reacting only on the external surface of crystalline zeolites. By varying the size of the OSDA for synthesizing the EZ, materials with controlled porosity, acidity, accessibility, and catalytic activity are prepared and their properties extend those of existing crystalline zeolites.

INTRODUCTION

Molecular sieve zeolites are widely used as heterogeneous catalysts in a variety of industrial (oil refining, petro- and fine chemicals production...) and environmental (exhaust gas treatment, heavy and radioactive ion sequestration, ...) applications.¹⁻² They typically offer high surface area, uniform micropores with well-defined topology, strong acid sites, and high thermal and chemical stability.³⁻⁵ A unique property of zeolite materials is their shape selectivity,³ a consequence of their active sites being confined in well-defined micropore spaces. However, the presence of such micropores often imposes transport limitations on reactants and products, in particular, those with a size close to or larger than the pore diameter.⁵ These steps become rate determining in the catalytic activity and the full potential of these catalysts is not reached. In the case of bulky molecules, larger than zeolite pores, the only accessible and active sites are those located on the outer surface of the crystals and their pore openings.

In summary, the problems of diffusion and related low accessibility to the active sites represent a major roadblock in a number of industrial reactions catalyzed by zeolites (cracking, oxidation, esterification, ...). To meet exacting industrial requirements,⁶ much efforts are devoted to preparing zeolite materials with

better accessibility and improved transport properties. The most explored avenues to improve catalytic performances were focused on decreasing zeolite crystal size⁷⁻¹¹ and preparing zeolites containing a system of larger (meso-)pores.¹²⁻¹⁷ A mesoporous network connected to the native micropores in zeolite crystals decreases mass transport limitations, but it does not affect the limited intrinsic ability of zeolites to process bulky molecules. Both solutions, nano-zeolites and mesopore-containing zeolites, bear some limitations, in particular the active sites in the zeolite channels remain inaccessible for bulky molecules, and therefore alternative solutions are required.

Recently, we addressed this issue by the preparation of zeolites units with sizes of few nanometers.¹⁸ These units denoted "embryonic zeolites" exhibit only short-range order and comprise elements of zeolite structure. Their semi-formed unit cells offer easier access of much bulkier molecules to the active sites than their fully crystalline counterpart, which exhibits long range order. The embryonic zeolites were prepared as mono disperse suspension containing nano particles that are uniform in size and composition.¹⁸ They exhibit superior catalytic activity, with respect to the fully crystalline material (zeolite) synthesized from the same initial system.

For instance, embryonic zeolites synthesized with tetrapropylammonium as structure directing agent exhibit pores in the

range 1-2 nm. The extension of zeolite pore size to 1-2 nm would allow to process much bulkier molecules than the “current” available zeolites can process. At present, only Ge-silicate zeolites exhibit pores larger than 1 nm,¹⁹ while the conventional aluminosilicate zeolites are limited to 1 nm in pore size. Unfortunately, the limited (hydro)thermal stability of Ge-silicate zeolites so far precluded their commercial use. Thus, the quest for extra-large pore zeolitic materials continues to be a topic of high priority. Here, we address this issue by the preparation of a few nanometer large zeolitic nano units which offer full accessibility to their active sites by relaxing diffusion limitations.

Sacrificial organic templates are commonly used in the synthesis of nanoporous materials.²⁰⁻²¹ In the synthesis of zeolitic materials the organic template, often called organic structure directing agent (OSDA) plays a key role in the formation of zeolite structure and charge distribution in the framework.¹⁹ In general, these are positively charged molecules that balance the negative charge introduced by Al cations in the zeolite framework. Each OSDA has a fixed size/charge ratio; therefore, the total guest volume is determined by the number of guest molecules that will correspond to the framework charge density.²²⁻²⁴

The objective of this study is the preparation of embryonic zeolites with pores exceeding 1 nm with a few nanometer particle size and controlled pore size and pore volume. A series of tetraalkylammonium (TAA) cations with size ranging between 6 and 12 Å was employed. The series includes: tetramethylammonium (TMA⁺), tetrapropylammonium (TPA⁺), trimethyladamantylammonium (TMAd⁺), tetrabutylammonium (TBA⁺) and tetrahexylammonium (THA⁺). With the exception of THA⁺, which is too big to yield a zeolite-type microporous material, all other OSDA generate zeolite structures.²⁵ Embryonic zeolites with different micropore volumes and specific surface areas are obtained. Their catalytic activities are evaluated in the dealkylation of a model molecule, 1,3,5-triisopropylbenzene (1,3,5-TiPBz) hardly penetrating the micropores of any industrially relevant zeolite. It is, therefore, commonly used to study the external surface properties of zeolites, including those with the large pores.^{26,27}

EXPERIMENTAL SECTION

Preparation of embryonic precursors

All syntheses started from clear homogenous sols with molar composition 9OSDA : 0.25Al₂O₃ : 25SiO₂ : 430H₂O : 100EtOH, where OSDA: Tetramethylammonium hydroxide (TMA.OH 25% in water, Alfa Aesar), Tetrapropylammonium hydroxide (TPA.OH 20% in water, Alfa Aesar) Trimethyladamantylammonium hydroxide (TMAd.OH 25% in water, Sachem), Tetrabutylammonium hydroxide (TBA.OH 40% in water, Alfa Aesar) and Tetrahexylammonium hydroxide (THA.OH 20% in water, Sigma Aldrich). After mixing the OSDA with distilled water, aluminum sulfate (98%) and tetraethylorthosilicate (TEOS 98%), the solution was hydrolyzed at room temperature for 6 h in a closed polypropylene bottle. The complete hydrolysis was confirmed by the appearance of a clear monophasic solution. Then, this solution was either dried by freeze drying at -94°C under vacuum (to avoid any further crystallization) or hydrothermally treated at 90°C for 12, 15, 18, 144 and 216 h as the kinetics of the synthesis depends on the nature of the OSDA. After the synthesis, the solid was recovered by freeze drying and further calcined at 550°C for 5h in a static oven under air.

The following series of zeolite precursor (P) samples were prepared by:

- Room temperature synthesis (P_{RT}): P_{RT}-TMA, P_{RT}-TPA, P_{RT}-TBA, P_{RT}-TMAd, and P_{RT}-THA.

- Hydrothermal synthesis: P_{x-y}-SDA, where x is the temperature in °C; and y the crystallization time in h; ex: P₉₀₋₁₅-TMA, P₉₀₋₂₁₆-TMAd, ...

A highly crystalline nano-sized ZSM-5 prepared from the same initial suspension at 100 °C for 48 h was used as a reference sample.

Characterization

All materials were characterized by powder X-ray diffraction (PXRD) using a PANalytical X'Pert Pro diffractometer with average Cu K α radiation ($\lambda = 1.5418$ Å). We used a θ -2 θ scan in the 3-50° 2 θ -range and a 0.02° step. The whole pattern was fitted using the combined analysis formalism²⁸ implemented in the MAUD program²⁹ a software using an extended Rietveld³⁰ analysis approach. Since severe line broadening was observed in the diffraction diagrams, we analyzed those within the Combined Analysis approach. The instrumental contribution to the line broadening was calibrated on the LaB6 srm660b standard powder from the National Institute of Standards and Technology. The Popa³¹ formalism was then used to describe anisotropic crystallite sizes and shapes. No preferred orientation was detected during the fit. The ZSM5 structural model n°1505105 from the Crystallography Open Database³² was used as a starting model in the refinement.

Nitrogen adsorption measurements were carried out on a Micromeritics ASAP 2020 surface area analyzer. The calcined samples were analyzed after degassing at 300 °C. The isotherms were recorded using the ASAP 2020 analysis program. The microporous volume ($V_{mic} / \text{cm}^3 \cdot \text{g}^{-1}$) and the external surface area ($S_{ext} / \text{m}^2 \cdot \text{g}^{-1}$) were obtained from the t-plot method based on the Harkins-Jura equation. The macro- and mesopores size distribution was obtained from the desorption branch using the Barrett Joyner Halenda algorithm assuming cylindrical pores. For refined analysis, micro-mesopores size distribution was extracted from Density Functional Theory modelling of the adsorption branch. TG-DSC measurements of different samples were carried out on a SETSYS evolution instrument (SETARAM). About 10 mg of each sample was introduced in an alumina crucible that is loaded in the analyzer chamber. The sample was heated from 30 °C to 800 °C with a heating ramp of 5 °C·min⁻¹ under air (flow rate: 40 mL·min⁻¹).

All NMR measurements were done with 4-OD mm zirconia rotors and with a spinning speed of 12 kHz. The solid state ²⁹Al CP-MAS NMR spectra were recorded on a Bruker Avance 400 spectrometer operating at 100.6 MHz. During the pulse program, a $\pi/2$ pulse of 3.7 μs on ¹H, a contact time of 2 ms and a recycle delay of 2 s were used. ²⁷Al MAS NMR was recorded on a Bruker Avance 500 spectrometer operating at 130.3 MHz. ²⁷Al MAS NMR was obtained with a $\pi/12$ pulse and a recycle delay of 1 s. The amount of nontetrahedral Al was estimated by the integration of peak area using Dm-fit program. The results were verified by the Top Spin program. The solid state ²⁹Si NMR (1pulse) was also recorded on Bruker Avance III-HD 500 spectrometer (11.7 T) operating at 99.3 MHz, using zirconia rotors of 4 mm outer diameter spun at 12 KHZ. A single pulse excitation (30° flip angle) is used with a recycle delay of 30s. TMS was the reference for ²⁹Si and ¹³C while a 1M Al (NO₃)₃ solution was used for ²⁷Al.

^{129}Xe NMR experiments were performed on a Bruker AV III 400 wide-bore spectrometer equipped with a 10 mm BBO probe, operating at a frequency of 110.64 MHz for ^{129}Xe . Prior to each experiment, the sample was introduced into a home designed (10 mm O.D.) NMR tube and dehydrated under high vacuum overnight at 413 K. The hyperpolarized (HP) ^{129}Xe gas was obtained using a home-built xenon polarizer based on the spin-exchange optical pumping (SEOP) technique.³³⁻³⁴ HP ^{129}Xe NMR spectra were acquired under a continuous recirculating flow of 50 mL·min⁻¹ of a gas mixture containing 4% Xe, 6% N₂ and 90% He at a total absolute pressure of 1.5 bar. For all variable-temperature (VT) measurements and after each temperature change performed at a speed of 5 K·min⁻¹, a delay of ~20 min was generally necessary to ensure a homogeneous temperature all over the sample, and this stability was checked by directly monitoring the xenon chemical shifts. All spectra were referenced to free xenon gas at 0 ppm. The IR spectra were recorded on a SHIMADZU IR Affinity-1 Fourier Transform Infrared Spectrophotometer using the KBr pellet technique. Pyridine and collidine adsorption were used to evaluate the acidity and accessibility of the samples. The study was performed using a standard in-situ FTIR set-up. Infrared spectra were recorded on a Nicolet Magna 550 FTIR spectrometer equipped with a DTGS detector at 4 cm⁻¹ optical resolution, with one level of zero-filling for the Fourier transform. Prior to the measurement, the sample was grinded and pressed into a self-supporting disc (diameter: 2 cm, approx. 5 mg cm⁻²) and activated in vacuum (ca. 10⁻⁶ hPa) at 460 °C for 2 h at 2 °C·min⁻¹. After cooling to room temperature, the spectrum of the sample was recorded for further use as a reference. Then, a pressure of 1.33 hPa of probe molecule (P.M.) was established in the cell at ambient temperature to reach saturation. The wafer was heated at 100 °C for 15 min to facilitate diffusion of P.M. into the sample. Successive evacuations were performed at 25 °C, 50 °C, 100 °C, 150 °C, 200 °C, 250 °C, 300 °C, and 350 °C at 15 min intervals. All spectra were normalized to 2.5 mg cm⁻² wafer. The amounts of acidic sites were determined using the areas of the bands at 1450 cm⁻¹ (Lewis) and 1545 cm⁻¹ (Brønsted) in case of pyridine adsorption and at 1644 cm⁻¹ (Brønsted) in case of collidine-adsorption. The molar extinction coefficients (ϵ) used for quantification were ϵ_{1545} (B-pyridine) = 1.8 cm²·mol⁻¹, ϵ_{1455} (L-pyridine) = 1.5 cm²·mol⁻¹ and ϵ_{1644} (B-collidine) = 10.1 cm²·mol⁻¹.³⁵ The OMNIC version 7.3 SP1 program was used for data processing.

The Transmission Electron Microcopy (TEM) investigations were carried out on a double corrected Cold FEG Jeol ARM Analytical TEM. Due to the materials high sensitivity to electron beam irradiation, an accelerated voltage of at 80kV was employed and the micrographs were acquired under Scanning TEM (STEM) mode using the High Angular Annular Dark Field (HAADF) detection methodology. In this approach, a 0.1 nm diameter beam scanned the region of interest and the image was recorded at a camera length of 8 cm and in the semi angular range 70 – 280 mrad. Size calibrated Au NP were employed for the STEM fine alignments. The images were acquired at a speed of 12 $\mu\text{s}/\text{px}$ for image sizes of 1024x1024 px using the Digiscan Plugin implemented in the Digital Micrograph software. Few drops of an ethanol-based solution containing the specimens dispersed in ethanol and using ultrasounds were deposited on a holey-carbon membrane disposed on 300 mesh Cu-grids.

Catalytic testing

The catalytic performances were evaluated in the dealkylation of 1,3,5-triisopropylbenzene (TiPBz) in a down flow atmospheric reactor. In a typical catalytic test, 20 mg of catalyst (dry mass) was loaded at the center of a stainless-steel tubular reactor (internal diameter of 1/2") and activated in-situ at 460 °C (ramping from room temperature at 5 °C·min⁻¹) under a dry air flow (50 ml·min⁻¹) for 1 h. The temperature was subsequently lowered to 300 °C under nitrogen flow (50 ml·min⁻¹). The nitrogen flow was then diverted to a saturator (operated at 70 °C) filled with TiPBz (Alfa Aesar, 97%), generating a TiPBz partial pressure of 170 Pa feeding the reactor. The WHSV (Weight Hourly Space Velocity) was held constant at 8 h⁻¹ for all tests. The conversions were measured at 300 °C over 100 min. The online analysis of the products and unconverted reactant, transferred via a line heated at 150 °C to a gas sampling valve, were monitored by a Varian CP-3800 Gas Chromatograph equipped with a flame ionization detector (FID) detector and a HP-PONA cross-linked methyl siloxane column, 50 m (L) x 0.2 mm (ID) x 0.5 μm (film thickness). The temperature of the column was kept at 160 °C throughout the measurement. The injection temperature was maintained at 200 °C while the detection temperature was maintained at 250 °C.

RESULTS AND DISCUSSION

Discrete nanometer sized embryonic zeolites can only be obtained if the polymerization between silica and alumina precursor species and the OSDA is closely controlled, to avoid the formation of large and dense gel particles with various sizes and compositions. This is also a necessary condition to assess their physicochemical properties. Such a general rule is particularly important with amorphous material lacking long range structural order. Optically clear sols containing only discrete gel particles are therefore used in all our syntheses to harvest solids with similar size, composition, and porosity.

Impact of OSDA on the physicochemical properties of embryonic zeolites.

Alkylammonium cations of varying sizes acting as organic structure directing agents (OSDAs) allow studying the properties of the porous solids harvested from syntheses performed at room temperature (RT). After complete hydrolysis of the silica source (TEOS), the resulting samples are freeze dried and characterized.

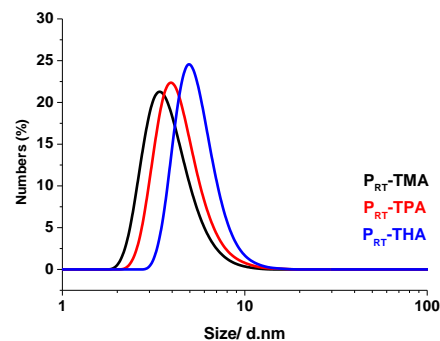


Figure 1. The particle size distribution of P_{RT}-TMA, P_{RT}-TPA and P_{RT}-THA suspensions.

According to DLS, all precursor sols present a monomodal particular size distribution with a size between 1 and 10 nm, **Figure 1.** The size of embryonic zeolites depends on the size of the template employed and a clear relation is observed: maxima in the DLS peak of P_{RT}-TMA, P_{RT}-TPA, and P_{RT}-

THA are 2 nm, 3 nm and 4.5 nm respectively. The synthesis conditions are such that no material exhibits any long-range structural order, monitored by X-ray diffraction, **Figure S1**. An SEM inspection highlights a uniform mass of aggregated particles, **Figure S2**, at higher magnification, the grain-like structure of the particles building the aggregates appears. More information on the size and morphology of the embryon-

ic zeolites was obtained by STEM-HAADF. Upon freeze-drying, discrete particles are observed. Their size (3-5nm) matches well with the result of DLS analysis. The particles generally exhibit a spherical morphology with a rough surface most probably due to the presence of pores with diameters <0.9 nm as indicated by the arrows in the inset of **Figure 2B**.

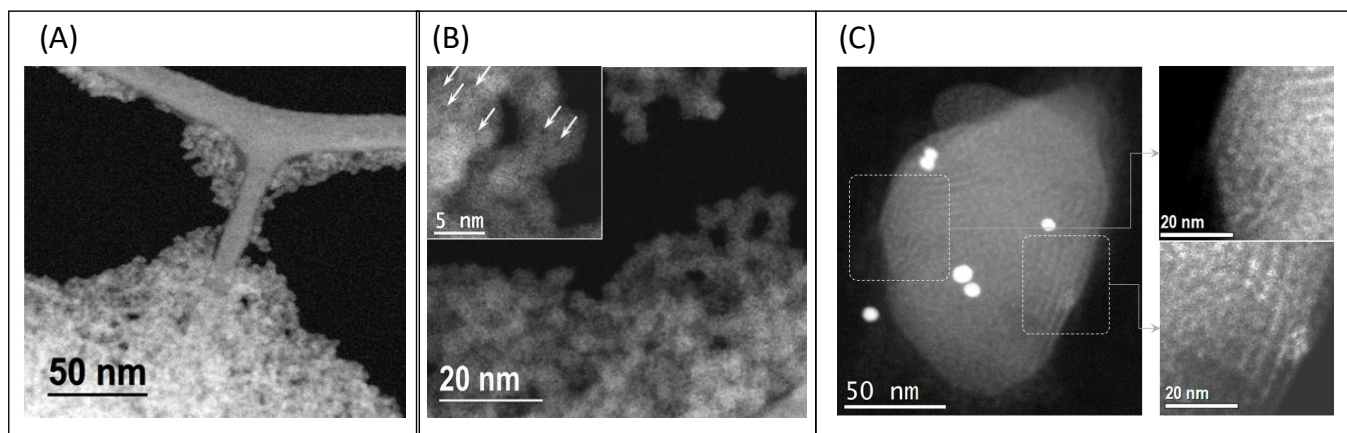


Figure 2. Low (A) and high (B) magnification view of freeze-dried PRT-TPA sample and colloidal crystal (C left) made of uniform in size PRT-TPA zeolite embryos and a high magnification view of the arrays of EZ building the colloidal crystals (C right). Note: the white spots in the left image are gold particles employed for the alignments.

Upon slow evaporation at room temperature, colloidal formation with arrays of individual EZ similar in size can be identified, **Figure 2C**. The formation of a crystal-like entity from the EZ shows the uniform nature of the colloidal suspen-

sion obtained after the RT synthesis, where only discrete particles of similar size are present. Besides, the agglomerate exhibit area of short-range order with the spherical EZ as units.

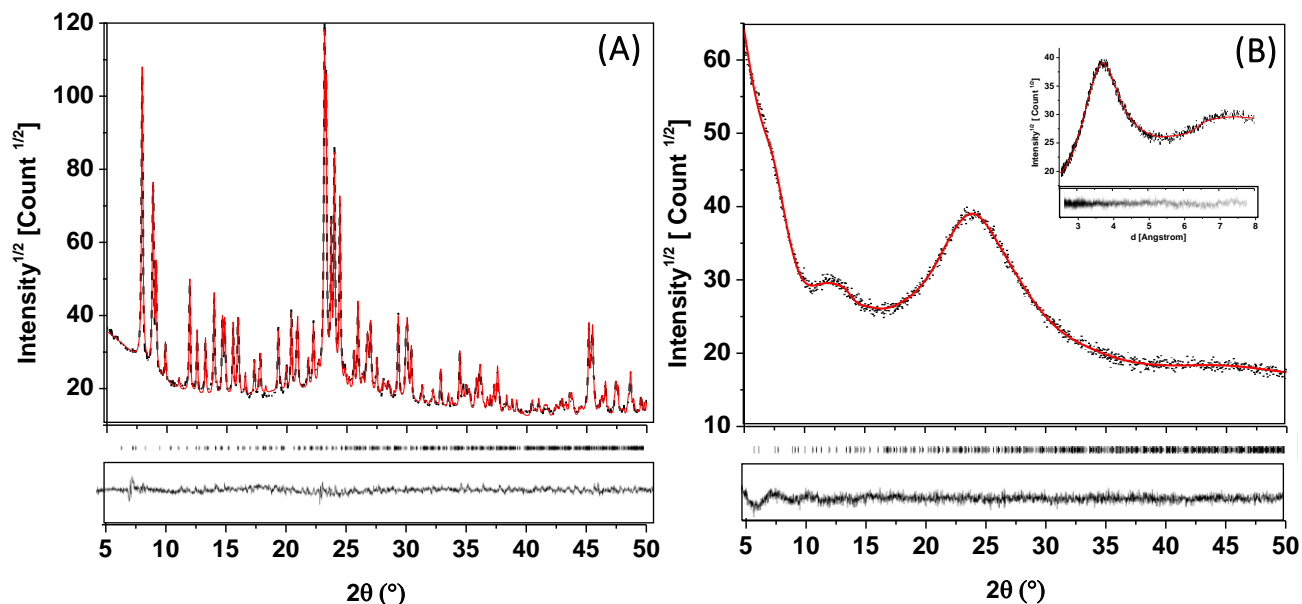


Figure 3. XRD diagrams and Rietveld fits for (A) well crystallized (mean crystallite sizes are 133(10) nm; GoF = 1.9) and (B) embryonic samples (GoF = 1.2). Diagram bottoms are difference curves, insert in (B) is plotted in d-scale.

Figure 3 shows the difference in XRD diagrams between fully crystallized and amorphous EZ prepared with TPA. While the fully crystallized powder (see **Figure 3A**) clearly exhibits narrow diffraction peaks with cell parameters coherent with the original structure ($a = 2.00484(7)$ nm; $b = 1.99267(6)$ nm; $c = 1.33919(5)$ nm), the amorphous sample (see **Figure 3B**)

only gives rise to first- and second-order modulations due to closest and second closer neighboring atoms of the structure. We could fit this latter diagram using crystallite sizes typically of the order of the unit-cell (typically few nm), and significant microstrains (typically 0.05 rms), with slightly isotropized cell parameters ($a = 1.906(6)$ nm; $b = 1.70(2)$ nm; $c = 1.55(1)$ nm).

Such fitted parameters are only helping the fit, but could not represent the first significant contribution appearing below $2\theta = 10^\circ$. This latter contribution is associated with the periodical arrangement of the embryos depicted by the electron density contrasts observed in TEM images (see **Figure 2C**) and cannot be modeled in the XRD diagrams using the MFI-type structure. We instead used a simple Gaussian background contribution to account for it in the Rietveld fits. As a result (see **Figure 3B inset**), the first and second scattering contributions are located around 3.68 and 7.1 Angstroms, respectively. In MFI-type structure, such contributions characterize the mean distances between closest Si-O polyhedra, with corresponding first and second scattering orders. The first distance at 3.68 Angstroms is also close to the one observed in fumed silica of around 3.9 Å (see **Figure S3A**) while no periodicity is observed (see **Figure S3B**).

As X-ray diffraction is not the best technique to probe short-range orders, on such embryonic zeolites we resorted to infrared spectroscopy (IR) to characterize the structure of ultra-small zeolite embryos. The IR spectra display three main absorption bands around 1080 (with a shoulder at 1200 cm^{-1}), 800 and 460 cm^{-1} , **Figure S4**. The first two correspond to asymmetric and symmetric Si-O stretching motions and the last one to a bending Si-O-Si mode. A shoulder at 560 cm^{-1} is also observed and assigned to the bending vibrations of external Si-O.³⁶ The bands around 560 and 1200 cm^{-1} are generally attributed to the double-membered ring and external asymmetric stretching in the MFI structure, respectively³⁷ while, the optical density ratio of 560 and 460 cm^{-1} bands is commonly used to determine the zeolite crystallinity.³⁷ The crystallinities of the P_{RT}-TMA, P_{RT}-TPA, and P_{RT}-THA samples are similar according to such an IR study. The synthesized materials exhibit a type I isotherm with a significant uptake at low relative pressure followed by horizontal adsorption-desorption branches. A closer look at the isotherms reveals small differences between the material synthesized with TMA (P_{RT}-TMA) and larger templates. The isotherm of P_{RT}-TMA is type Ia characteristic of ultra-micropores (< 0.7 nm), whereas the others (TPA, TBA, TMA_d, THA) are of type Ib, characteristic of super-micropores (1-2.5 nm) according to the classification of IUPAC.³⁸ The latter isotherms show much higher uptake indicative of larger micropore volumes. The isotherm of nanosized ZSM-5 is included in **Figure 4** and the data derived from the isotherm summarized in **Table S1**.

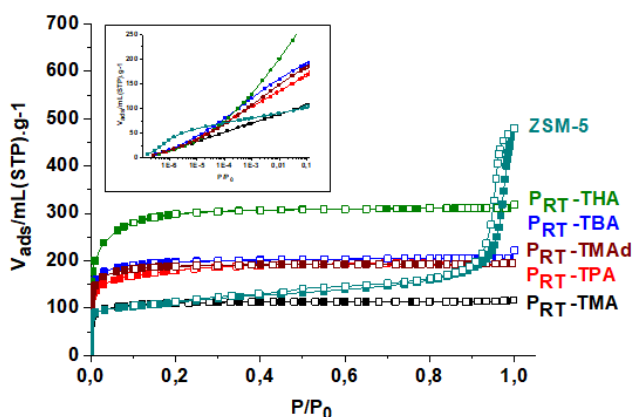


Figure 4. N₂ adsorption/desorption isotherms of P_{RT}-SDA samples, where SDA is TMA, TPA, TMA_d, TBA, and THA. The isotherm of nanosized ZSM-5 is included as a benchmark. The

inset is plotted in the logarithmic x-axis scale and represents the low (<0.1 P/P₀) pressure range.

The micropore volume of the MFI type material is similar to P_{RT}-TMA and much lower than that of the embryos synthesized with larger templates. However, the total pore volume of the crystalline material is higher due to textural meso-porosity, revealed by the slope of the isotherm in the 0.2 – 0.8 pressure range and the second uptake in the 0.9-1.0 range, between nanosized zeolite particles. Such a textural porosity is not present in the embryonic zeolite. If some textural pores exist, i.e. pores between ultra-small nano particles (3-4 nm) formed around the template, these pores remain in the microporous range.

The data summarized in **Table S1** show that the microporous volume and specific surface area increase with the size of the SDA used. More precisely, V_{micro} and S_{BET} increase from 0.17 to 0.47 $\text{cm}^3.\text{g}^{-1}$ and from 418 to 1095 $\text{m}^2.\text{g}^{-1}$ for TMA and THA, respectively. A small increase in micropore volume (TBA>TMA_d>TPA) is noticeable.

In general, templates with approximately similar sizes yield materials with very close specific surface area and micropore volume; a close positive correlation between the size of the OSDA and the V_{micro} and S_{BET} is indeed observed on these samples. It is well known that tetraalkylammonium cations play a key role in zeolite synthesis by ordering around them aluminosilicate species, akin to water clathrates, by progressive displacement of hydrogen bound water molecules.³⁹⁻⁴² As long as synthesis is stopped before crystallization, the harvested materials are silica-alumina matrices organized around these OSDAs and form cages of different volumes depending on the size of SDA used.

The pore size distributions and the cumulative pore volumes of P_{RT}-TPA and P_{RT}-THA, extracted by DFT using a cylindrical geometry (model “N₂-cylindrical pores-oxide surface”) are displayed in **Figure S5**. The EZ prepared with TPA has a wide distribution of pore widths ranging between 8 and 20 Å with the highest population of pores smaller than 14 Å; they correspond to a cumulative pore volume of 0.15 $\text{cm}^3.\text{g}^{-1}$. Moreover, the EZ prepared in the presence of THA does not contain pores with size below 1 nm. The population of large pores (> 14 Å) is the most largely presented in this sample. Despite their higher V_{micro} and S_{BET}, no additional larger pores are observed, so pores larger than 14 Å dominate in this material. This is clear from the increase of their incremental volume and the decrease of the cumulative pore volume of pores smaller than 14 Å from 0.15 to 0.09 $\text{cm}^3.\text{g}^{-1}$. Some textural pores represented by a large peak centered at about 27 Å, present in P_{RT}-TPA. These data unambiguously imply that the size of the OSDA determines the pore size of the embryonic zeolites; a smaller OSDA yields pores of smaller width while larger OSDA leads to larger pores.

The pore structure of EZ can be studied by relating the confinement of Xe molecule to their chemical shift.⁴³ Low chemical shift peaks, close to 0 ppm, indicate large (meso-)pores, while high chemical shifts indicate more confined space (micro) pores. The NMR spectra of HP ¹²⁹Xe adsorbed on the P_{RT}-TPA and P_{RT}-THA samples recorded at different temperatures are shown in **Figure 5**. In addition to the peak of xenon in the gas phase at 0 ppm, all spectra exhibit an additional resonance, which differs in P_{RT}-TPA and P_{RT}-THA samples. The P_{RT}-TPA material exhibits a single peak (136–92 ppm) in the 200–320

K temperature range. The peak in P_{RT}-THA sample shifts from 115 to 84 ppm in the 200–320 K temperature range. The higher chemical shift of Xe adsorbed in P_{RT}-TPA can be attributed to xenon atoms located in a more confined environment. This

result is in good agreement with the physisorption measurements showing that the large THA template generated EZ with substantially higher pore volume.

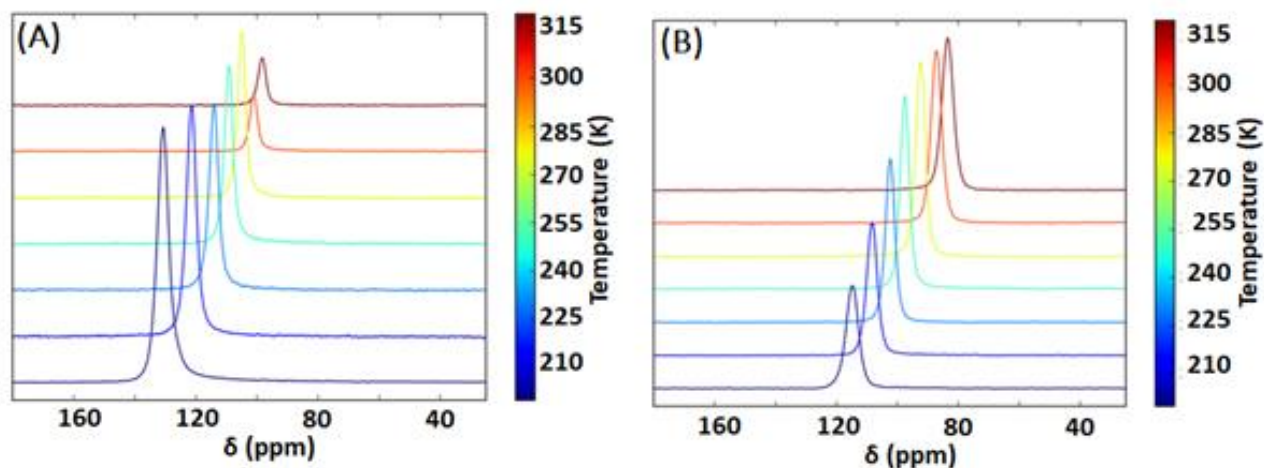


Figure 5. NMR Spectra of HP ¹²⁹Xe adsorbed on P_{RT}-TPA (A) and P_{RT}-THA (B) samples recorded at variable temperature (200-320K).

In order to gain more insight into the structure and properties of the embryonic zeolites, the samples are further characterized by thermal analysis (TG/DSC, **Figure S6**) and ¹³C CP MAS NMR spectroscopy, **Figure S7**.⁴⁴ Prior to TG-analyses, the samples are equilibrated overnight under a 79 % relative humidity at room temperature. The thermograms of different samples are fairly similar to two well established regions, a first, between room temperature and 200-250°C and a second between 250 and 500°C. The weight loss below 200°C is associated with an endothermic peak and attributed to water release. The weight loss at higher temperature (250-500°C) is associated with an exotherm (DSC) and corresponds to the OSDA oxidation. Compared to a crystalline zeolitic material, the embryonic zeolites are much more hydrophilic. For instance, the water content of ZSM-5 is about 5%, while its amorphous counterpart synthesized with TPA contains about 35 wt% water. It is of note that water content increases with the increasing size of the OSDA used, from 38 to 48 and 52% with TMA, TPA, and THA, respectively. Most of the water is released at about 60°C on P_{RT}-THA and at higher temperatures on samples containing smaller structure directing agents (P_{RT}-TPA and P_{RT}-TMA); this further implies that larger (hydrated) OSDAs lead to embryos with more open structures. The OSDA oxidation includes several steps; one common around 280-300°C, characteristic of weakly bound cations located on the external surface of the embryonic units, and another at higher temperature related to the size of OSDA. With a smaller OSDA, its decomposition is slower due to its confinement and more restricted access of the oxidant (O₂) highlighted by a shift to higher temperatures (450°C in PTMA vs. 350°C in PTPA/THA).

¹³C CP MAS NMR spectroscopy of the embryonic zeolites shows for each carbon frequency a superposition of a narrow signal and a large signal with a slight chemical shift. The narrow signals correspond to OSDA in less restricted environments, most probably located on the external surface of zeolite embryos, while the larger, correspond to OSDA in more confined environments, OSDA occluded into zeolite embryos, **Figure S7**. This is indicative of OSDA in two dif-

ferent environments.⁴⁵ The first sharp one belongs to mobile OSDA, most probably located on the external surface of zeolite embryos, while the second, broader, corresponds to OSDA located in a more confined space. This is consistent with the previously observed two exotherms on the DSC, namely: i) 200-250°C, OSDA in a loosely confined environment, ii) 250-500°C, OSDA in a more confined environment. Moreover, the ¹³C CP MAS NMR of TPA⁺ occluded in crystalline ZSM-5 has the distinctive splitting of its C3 (terminal methyl) peak attributed to the presence of two different environments (linear and zig-zag channels) in the zeolite channels.⁴⁶⁻⁴⁷ Such a splitting is not observed in P_{RT}-TPA, indicating the presence of its methyl groups (C3) in a unique environment. A shift of -3 ppm of the C1 peak in the EZ indicates a more distorted angular configuration of its methylene group attached to the nitrogen as reported earlier.⁴⁸

Figure S8A shows the ²⁷Al NMR spectra of all as-prepared samples, with a single peak centered at 53 ppm characteristic of tetrahedral aluminum. After calcination, **Figure S8B**, aluminum is present in other coordination: a majority remains in tetrahedral coordination (ca. 53 ppm), some become octahedral (0 ppm) and a broad peak (0 and -40 ppm) appears often attributed to distorted tetrahedral Al and/or penta-coordinated Al can be observed. The proportion of non-tetrahedral Al varies in the different samples. By deconvoluting the spectra using the dm-fit program, the relative fraction of octahedral aluminum is estimated between 38% (P_{RT}-TMA) and 26% (P_{RT}-THA), higher than the 10% usually observed for a pentasil type zeolite. However, this amount is substantially lower than what is observed in a variety of amorphous aluminosilicates.⁴⁹ This indicates that a substantial portion of Al is stabilized in a silica matrix and remains there upon calcination at high temperature. The dealumination of zeolites depends on both the framework type and the T site within a given zeolite.⁵⁰⁻⁵¹ The larger proportion of octahedral Al is observed in the calcined P_{RT}-TMA embryos and the lowest in P_{RT}-THA; this could be related to the bond angles of the alumina-silica species surrounding the template, as the small TMA generates

an inorganic network with more stress than the larger templates.

The acidity of embryonic samples is characterized by *in-situ* infrared spectroscopy using pyridine as a probe molecule. The

number of Brønsted acid sites (BAS) is calculated from the pyridinium band at 1545 cm^{-1} after the evacuation of pyridine at different desorption temperature. The results are given in **Figure 6**.

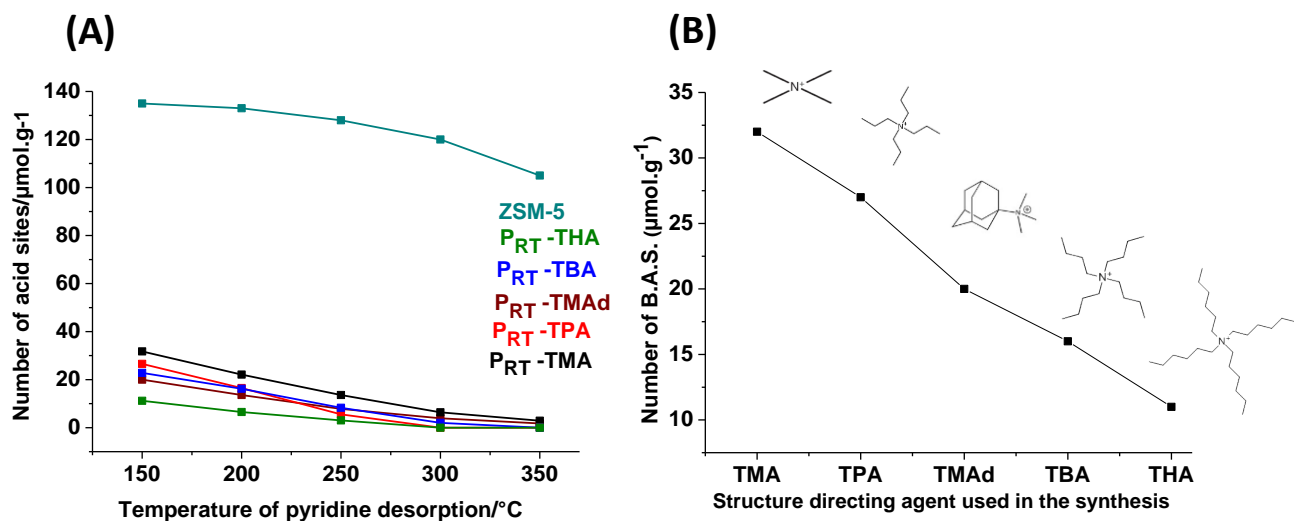


Figure 6. The number of Brønsted acid sites ($\mu\text{mol.g}^{-1}$) for PRT-OSDA and nanosized ZSM-5 at different temperatures determined by TPD of pyridine (A) and at 150°C as a function of OSDA size (B).

The embryonic zeolites possess a significantly lower concentration ($< 40\ \mu\text{mol.g}^{-1}$) and strength of Brønsted acid sites in comparison to ZSM-5 ($\sim 135\ \mu\text{mol.g}^{-1}$), attributed to a wider distribution of angular configurations and organization of the aluminosilicate matrix compared to highly crystalline zeolites. After pyridine desorption at 150°C , the embryonic zeolites have slightly different concentrations of BAS, depending on the employed OSDA. Smaller OSDA leads to higher BAS concentrations (**Figure 6B**). An increase in the size/charge ratio of an OSDA leading to a decrease in the number of BAS was already reported by Barrer and Denny,⁵² Aeillo and Denny⁵³ and Kerr⁵⁴ for different zeolites. After desorption at a higher temperature ($> 250^{\circ}\text{C}$), the number of Brønsted acid sites is close to zero for all samples, indicative of the low acid strength of the embryonic zeolites. By combining ICP-analysis and ^{27}Al NMR results, it is possible to evaluate the amount of Al^{IV} in each sample and, therefore the number of expected BAS. The Pyridine method account, however, for only 20% of the expected amount of BAS. This might be due to the presence of BAS in a very confined environment inaccessible to pyridine or the remaining Al^{IV} present in the calcined samples does not completely correspond to BAS. Indeed, the peak centered at 53 ppm in ^{27}Al NMR is not symmetrical and presents an extension between 40 and 10 ppm, which can be attributed to highly distorted, tetrahedral, non-framework, aluminum-oxygen species. The chemical shift around 30 ppm is also consistent with penta-coordinated aluminum.⁵⁵⁻⁵⁶ In addition, during calcination, the tetrahedral symmetry of the lattice aluminum must be lost, and aluminum species of low symmetry formed. In a low symmetry environment, aluminum can become “invisible” due to important peak broadening of this quadrupolar nucleus.⁵⁷⁻⁵⁹

The accessibility to active sites (B.A.S.) is best studied by infrared spectroscopy of adsorbed probe molecules of different kinetic diameters, namely Pyridine (0.57 nm) and Collidine (0.74 nm).³⁵ The results of such experiments, performed under

similar conditions, are summarized in **Figure S9**. The number of B.A.S is calculated from the bands associated with protonated pyridine (1545 cm^{-1}) and collidine (1648 cm^{-1}) after their adsorption and evacuation of the physisorbed portion of the adsorbates. The accessibility to B.A.S. is defined as the ratio between the number of B.A.S. measured by collidine and pyridine. While, with the TMA OSDA (PRT-TMA), the accessibility is 0.37, i.e. the sites probed by collidine are less than half of those probed by pyridine ($13\ \mu\text{mol.g}^{-1}$ vs $36\ \mu\text{mol.g}^{-1}$), it becomes 1 when TPA is the OSDA (PRT-TPA). This indicates that all active sites are accessible to both molecules in the latter. So, in the presence of small OSDA, small pores are formed, restricting accessibility of their Brønsted acid sites to bulky molecules and leading to potential shape selective effects in embryonic zeolites.

^{29}Si NMR spectra of as-synthesized and calcined embryonic zeolites prepared in the presence of the smallest (TMA), medium (TPA) and largest (THA) OSDAs show different silicon environments Q^2 , Q^3 and Q^4 (**Figure S10**).⁶⁰ Their deconvolution with the dm-fit program gives the relative fraction of each species. After calcination, the chemical shifts are unchanged but the relative areas of Q^2 and Q^3 decrease while the Q^4 increase. These changes result from two simultaneous events upon heating: i) some tetrahedral Al changing coordination, ii) condensation of some silica species. The calcined samples have almost the same ratio of Q^3 and Q^4 species but the contribution of Al and OH groups to the Q^3 species differs between different samples.

Impact of temperature and synthesis time on the physicochemical properties of embryonic zeolites

In addition to room temperature synthesis, embryonic zeolites are prepared at 90°C during 12-216 h. As with the RT syntheses, all samples are freeze-dried before further processing and characterization.

The synthesis time is carefully controlled to obtain preformed units with only short-range order. Only the system containing TPA yields a crystalline phase (MFI-type) in syntheses longer than 24 h. The other systems are always amorphous after crystallization up to 216 h (**Figure S11**).

The ^{13}C CP MAS NMR of the TMA, TPA, and THA templates and the embryonic zeolites synthesized at RT and 90°C are displayed in **Figure S7**. All characteristic peaks of the template are present in the spectra of the embryonic zeolites. Peak broadening, an indication of the presence of the template in a confined environment, is observed.

Like the embryonic zeolites obtained at RT, the solids synthesized at 90°C exhibit a type I adsorption isotherm characteristic of a microporous material. In general, the materials synthesized at 90°C show an increase in nitrogen uptake with respect to their counterpart obtained at RT, **Figure S12**. This suggests a better structuring of these embryonic zeolites. The maximum micropore volume is achieved after different synthesis times for different templates. For instance, it is 12 h for a TPA synthesis and it decreases after 18 h, **Figure S12B**; this is related to the onset of crystallization of MFI. With TBA, there is no difference between 15 and 144 h, **Figure S12C**, while with TMA the micropore volume increases up to 216 h of hydrothermal synthesis. The single exception is with THA, where the 90°C synthesis shows a lower pore volume and surface area ($0.38\text{ cm}^3\cdot\text{g}^{-1}$ and $911\text{ cm}^2\cdot\text{g}^{-1}$) than its RT counterpart ($0.47\text{ cm}^3\cdot\text{g}^{-1}$ and $1095\text{ cm}^2\cdot\text{g}^{-1}$), **Figure S12E** and **Table S2**.

It is difficult to determine the reason for such a trend but could be related to the large size of the THA molecule and a more efficient condensation of the aluminosilicate species at a higher temperature.

The effect of synthesis temperature on the state of Al in the embryonic zeolites is investigated by ^{27}Al MAS NMR and IR monitored pyridine desorption. The ^{27}Al MAS NMR spectra of all as-synthesized samples exhibit a single peak (ca. 53 ppm) indicating the sole presence of tetrahedral aluminum, **Figure S13**. However, after calcination at 550°C, the intensity of this peak decreases, indicating that a portion of aluminum leaves its tetrahedral coordination. At the same time, two peaks emerge: a sharp and small at 0 ppm, typical for octahedrally coordinated aluminum, and a broad between 0 and -30 ppm, attributed to the presence of different polymorph of aluminum. Deconvolution of these spectra with the dm-fit program can evaluate the amount of nontetrahedral aluminum. The octahedral aluminum content decreases in the 90°C synthesis. The tetrahedral aluminum peak (53 ppm) is not symmetrical in all samples and tails between 40 and 10 ppm. As mentioned above, this can be attributed to the presence of highly distorted, tetrahedral, non-framework, aluminum-oxygen species or penta-coordinate aluminum.

IR spectroscopy of adsorbed pyridine measures the acidity of the embryonic zeolites. The number of Brønsted acid sites, calculated from the pyridinium ion band (1545 cm^{-1}), are summarized in **Table 1**.

Table 1. Number of Brønsted acid sites ($\mu\text{mol}\cdot\text{g}^{-1}$) of P-OSDA prepared at RT and at 90°C.

Embryonic samples	P-TMA		P-TPA			P-TMA _d		P-TBA		P-THA		ZSM-5	
	RT	P ₉₀₋₁₅	RT	P ₉₀₋₁₅	P ₉₀₋₁₈	RT	P ₉₀₋₁₅	P ₉₀₋₂₁₆	P ₉₀₋₁₅	P ₉₀₋₁₄₄	RT		P ₉₀₋₁₄₄
150	32	34	27	35	38	20	40	37	23	20	11	22	135
200	22	23	17	27	29	14	27	26	16	15	7	15	133
250	14	12	6	18	18	8	15	15	8	7	3	7	128
300	6	5	0	8	9	4	5	6	2	0	0	0	120
350	3	1	0	2	2	2	2	1	0	0	0	0	105

The acidity of the EZ prepared at 90°C increases with synthesis time, but differences are marginal. As for the room temperature EZ syntheses, the smaller the OSDA, the higher the number of BAS; differences, however, are not substantial and all EZ possess a much lower number of acid sites than ZSM-5 with similar Si/Al ratio.

Catalytic activity

Earlier studies showed that amorphous zeolite precursors compelling catalytic performances in hydrocarbon conversion.⁶¹⁻⁶² In the present work, we evaluated the catalytic performances of embryonic zeolites prepared at RT in the dealkylation of triisopropylbenzene (TiPBz), **Figure 7A**. Its kinetic diameter (0.95 nm), is well above the pore openings of the MFI-type zeolite (0.56 nm) and TiPBz is therefore com-

monly used to study the external surface properties of zeolites, including the large pores (12 MR) ones.

P_{RT}-TMA_d and P_{RT}-TMA EZ lead to the highest and the lowest conversion, respectively, **Figure 7A**. **Figure 7B** highlights the effects of the OSDA on the conversion, number of Brønsted acid sites and micropore volume. While P_{RT}-TMA possesses the highest number of Brønsted acid sites, its smallest conversion of TiPBz is related to the limited accessibility of its acid sites. This clearly indicates that the OSDA size is a tool to control the size selectivity of the EZ. For instance, TPA and TMA_d generated materials comprise relatively strong Brønsted acid sites in micropores large enough to process TiPBz. Conversely, the larger TBA and THA OSDA generate EZ with larger micropores accessible to bulky molecules, but their catalytic performances are low due to a lack of Brønsted acid sites. Thus, EZ combining a critical number of active sites

with pores opening large enough to transport the bulky TiPBz display the best catalytic performances; the catalytic properties

of embryonic zeolites can be tuned by accessibility and acid/active site density.

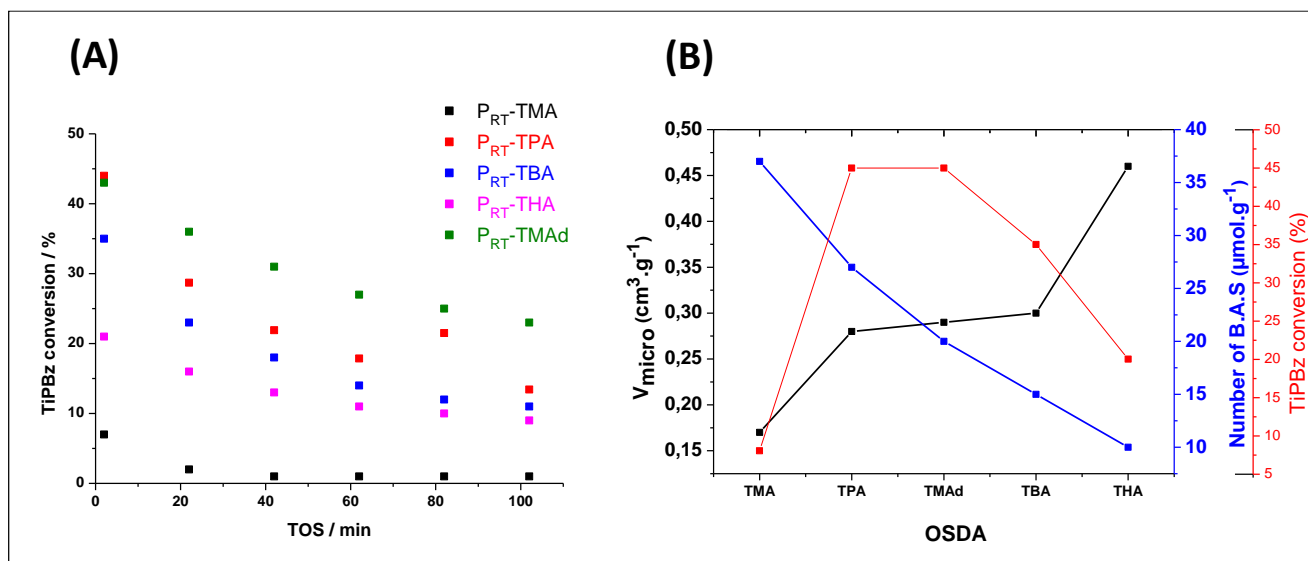


Figure 7. (A) TiPBz conversion of RT synthesized embryonic zeolites (A). (B) TiPBz conversion, number of Brønsted acid sites and micropore volume of the embryonic zeolites as a function of OSDA in the synthesis. Catalytic test conditions: T = 300°C, WHSV = 8h⁻¹, TOS = 2 min.

CONCLUSIONS

X-ray amorphous zeolite precursors, the so-called embryonic zeolites (EZ) are prepared with OSDA identical to those used in the synthesis of conventional zeolites. The synthesis conditions need to be controlled to obtain monodisperse suspensions of zeolite precursor particles with a size of a few nanometers. A thorough physicochemical characterization indicates that such zeolite precursors are highly uniform, and their properties strongly depend on the template used. All are microporous materials as the majority of their pores are in the range 1-2 nm. Some exhibit impressive specific surface area (1095 cm².g⁻¹) and micropore volume (0.47 cm³.g⁻¹), larger than that of zeolites obtained with the same template. They are stable upon calcination and can be used as heterogeneous catalysts. While they possess a relatively low number of active sites of moderate acid strength, they display a high activity in bulky molecules conversion as their performance depends on and can be tuned by controlling their acidity and the accessibility of their active sites. EZ, therefore, extends the range of applications of current zeolites especially for bulky molecules; these are encountered in the processing of heavy oil fractions and biomass.

ASSOCIATED CONTENT

Supporting Information

The Supporting Information is available free of charge on the ACS Publications website at DOI:

Characterization data: XRD, SEM, FTIR, N₂ adsorption/desorption, Pore size distribution (DFT calculation), TG/DSC, NMR (pdf).

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Notes

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