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## Article

# Benefits of Ryegrass on Multicontaminated Soils Part 2: A Green Process to Provide Idrocilamide

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**Abstract:** A restoration of highly contaminated garden soil is proposed as a greener alternative to the production of vegetables. Depending on potentially toxic elements and their concentration, ryegrass shoots accumulate these elements in sufficient quantity to be used as a catalyst in organic synthesis. The analysis of Cd, Pb, Zn, Cu, Fe, Mn, and Al in ashes issued from the calcination of ryegrass shoots revealed that the concentration of Zn was highest ( $>7000 \text{ mg kg}^{-1}$ ). The ratios between potential Lewis acids (Zn, Cu, Fe, Mn, and Al) to carcinogenic metals (Cd or Pb) were  $191 \pm 7$  for LA/Cd and  $235 \pm 13$  for LA/Pb, making the shoots of ryegrass suitable for the production of Zn-rich polymetallic biosourced catalysts. This material was used in the synthesis of idrocilamide under free-solvent condition, providing the drug in a good yield (69.9%). Data show that a limitation of waste, a maximization of the material incorporation in the process, a minimization of the steps, and an optimization of the stoichiometric factor are the main innovative factors in the current process in comparison with those previously reported.

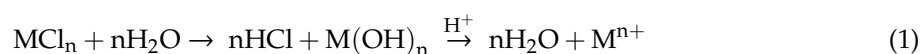
**Keywords:** ryegrass; green chemistry; biosourced catalyst; synthesis; metrics

## 1. Introduction

The problem of multi-contaminated sites by inorganic pollutants is a worldwide concern since the potential toxic elements (PTEs) are persistent and ubiquitous in the earth's environment. Among the 10 million sites that have been reported as being contaminated, 50% are contaminated by PTEs [1]. In addition to fertilizer application and atmospheric deposition, contaminated soils from urban areas (recreational areas, lawns, and kitchen gardens) are often subjected to other sources of contamination such as traffic emission, urban effluent, waste disposal, and the addition of ashes from wood- and coal-burning. These dangerous and widespread sources of contamination are of major concern since (i) it is recognized that the bioavailability of PTEs from anthropogenic contamination is higher than geogenic sources [2], (ii) kitchen gardens (shared, allotment, and private gardens) are often more contaminated than agricultural soils [3–5], (iii) homegrown vegetables, fruiting vegetables, and fresh

herbs do not comply with the European legislation due to the accumulation of PTEs [6–13], (iv) the health of people living in urban areas and the surrounding ecosystems are affected [14–16]. In this context, some gardeners have taken the decision to change their cultural practices whereas others decided to stop gardening and to sow ryegrass. This plant is described as a suitable plant for the revegetation of contaminated soils from metallurgical sites, able to remove Cd from soils, and to survive in contaminated acidic soils [17–20]. For these reasons, this study focuses on the possibility of finding a new valorization way for contaminated shoots of ryegrass.

Currently, people leave contaminated grass clipping in their own lawn as mulch and/or their own garden as compost or get them disposed. Although the preferred disposal route in EU countries is composting (improvement of microbiological, biochemical, physicochemical, and agronomical properties of soils [21–24]), some drawbacks in using contaminated composts have been highlighted [25–33]. Thus, to avoid the disposal of the harvested contaminated biomass, physical and/or chemical processes were developed after the incineration of this biomass, thus enabling the recycling of nutrients and/or PTEs. Three strategies, called agromining, phytomining, and ecocatalysis have been recently described as promising technologies in recovering PTEs from presently untapped secondary resources [34–39]. The use of polluted ashes as ecocatalysts seems to be very interesting since one industrial challenge is to replace metal-chlorides used in organic synthesis as Lewis acids ( $\text{MCl}_n$ ; Equation (1)). They are toxic for humans, react with water to produce  $\text{HCl}$  (Equation (1)), and they have to be removed from organic products, thus generating chemical waste. These ecocatalysts were described as new eco-friendly plant-based green Lewis acid catalysts [34–37] and may be considered as part of the progress in sustainable chemistry. This has become ever truer since Professors Anastas, Warner, and Zimmerman stated the 12 Principles of Green Chemistry and the 12 principles of green engineering in which this concept is embodied [40,41].



In the first part of this study [42], we demonstrated that the shoots of ryegrass may have a potential interest in organic chemistry. Based on our previous findings and considering the 12 principles of green chemistry and green engineering, a new experimental design was conducted and a new valorization way of contaminated ryegrass shoots was proposed. A biosourced catalyst was thus produced from contaminated ryegrass and was used to synthesize idrocilamide. This organic compound is the main ingredient of several drugs with anti-inflammatory and muscle relaxing properties (e.g., Srilane®).

## 2. Materials and Methods

### 2.1. Perennial Ryegrass Shoots Production

A pot (18 cm in diameter, 14 cm height) experiment was established in a greenhouse for 8 weeks using the soil G1 described in the first part of the current study [42]. Containers ( $n = 6$ ) with 2 kg of soil G1 were placed in the greenhouse with day/night temperatures of  $28 \pm 5 / 10 \pm 5$  °C (corresponding to April–July). Tap water (electrical conductivity,  $10 \text{ dS m}^{-1}$ ; pH, 7; Cd concentration,  $0.9 \text{ } \mu\text{g L}^{-1}$ , and Pb concentration,  $2.1 \text{ } \mu\text{g L}^{-1}$ ) was added in each pot to maintain the soil water content at 60% field capacity for two months. After this period, ryegrass seeds (*Lolium perenne* L.; 1.5 g) were sown, and the shoots were harvested 8 weeks after sowing. The total fresh shoot biomass was 408 g ( $68 \pm 7$  g per container;  $n = 6$ ).

### 2.2. Digestion and Determination of PTE Concentrations in the Biomass of Shoots

Air-dried plants were crushed with a knife crusher (Grindomix GM 200, Retsch) and sieved to pass  $315 \text{ } \mu\text{m}$ . Subsamples ( $0.3 \text{ g}$ ;  $n = 6$ ) were placed in a 50-mL digestion tube, and 5 mL of  $\text{HNO}_3$  (69%) was added. The mixture was firstly digested at  $65 \pm 5$  °C for 20 min, and then, at  $90 \pm 5$  °C for 45 min (HotBlock™ Environmental Express®SC100, Charleston, SC, USA). After cooling, 1.2 mL of deionized water and 1.8 mL of  $\text{H}_2\text{O}_2$  solution (30%) were slowly added, and the reaction medium was

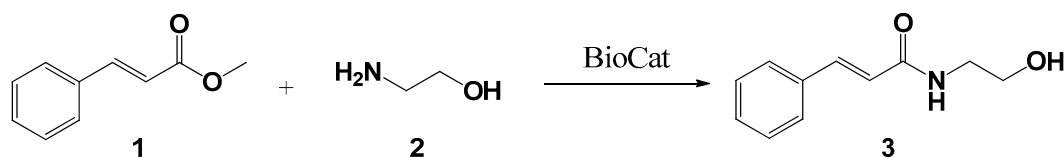
heated at  $90 \pm 5$  °C for 45 min. After cooling, the volume was adjusted to 25 mL with deionized water. Each digested-solution was filtered over an acetate Millipore membrane (0.45 µm porosity) and stored in acid-washed plastic bottles at 4 °C prior to analysis. The concentrations of PTEs in certified samples and shoots biomass were determined by absorption spectrometry (FAAS; AA-6800 Shimadzu, Tokyo, Japan) following the recommendation described in detail in the literature [43–45]. Limits of detection and quantification are detailed in the first part of the current study [42].

### 2.3. Preparation of Biosourced Catalyst

Air-dried shoots of ryegrass (70 g) were thermally treated in a muffle furnace (Nabertherm P330, Lilienthal, Germany) under airflow. Two warm-up phases were defined: from 20 to 250 °C in 30 min and stabilization for 1 h and from 250 to 500 °C in 2 h and stabilization for 8 h. After an overnight cooling period, the ashes were put in a flask, fitted with an air condenser, and a 2 M HCl solution was slowly added in a ratio 1/10. The mixture was stirred and heated at 70 °C for 2 h. After cooling, the solution and the residue were separated using celite, and the solution was concentrated to obtain a solid residue. After drying at 80 °C for 12 h, the yellow dry residue was mixed with montmorillonite K10 (MK10) in order to obtain a solid. The molar quantity of Zn in the solid was 1.22 mmol/g MK10. The solid mixture was stirred at room temperature for 5 h after adding methanol. In the last step, the solid was filtered and was thermally activated at 120 °C [46,47].

### 2.4. Synthesis of Idroclamide under Green Conditions

Methyl cinnamate (1; 0.5 g) and 2-aminoethanol (2; 0.19 g) were mixed together with the biosourced catalyst (BioCat; 126 mg) at 70 °C for 24 h under solvent-free conditions (Figure 1). After cooling to room temperature, the mixture was filtered and the compound was purified on flash chromatography (EtOAc:*n*-Heptane, 0:1 to 1:1). <sup>1</sup>H-NMR spectrum was run on a Varian 400-MR spectrometer with tetramethylsilane (TMS) as an internal standard at 25 °C in order to control the purity of idroclamide (3; data not shown).



**Figure 1.** Amidation of methyl cinnamate using the biosourced catalyst (Biocat).

### 2.5. Green Chemistry Metrics

The atom economy is a theoretical value described in the literature and expressed according to Equation (2) [48]. From a practical point of view, this metric is related to the percentage of reactants engaged in the reaction process that may be considered as waste materials.

$$AE = \frac{\text{Molecular weight of the isolated product}}{\text{Sum of the molecular weights of the reactants}} \times 100 \quad (2)$$

The E-factor is calculated following Equation (3) [49]. For researchers and industrials, this metric measures how much material enters the production site and how much leaves as product and/or waste materials.

$$E - \text{factor} = \frac{\text{Mass of waste}}{\text{Mass of product}} \quad (3)$$

The reaction mass efficiency is calculated from Equation (4) [50].

$$RME = \frac{\text{Isolated weight of the isolated product}}{\text{Sum of the weights of the reactants}} \quad (4)$$

The stoichiometric factor (Equation (5)) highlights the difference between the reaction conducted in the process studied and this reaction when the reactants are used in a stoichiometric condition [50].

$$SF = 1 - \frac{\text{Sum of the weights of the reactants} - \text{sum of the weights of the reactants in a stoichiometric process}}{\text{Sum of the weights of the reactants in a stoichiometric process}} \quad (5)$$

### 3. Results and Discussion

#### 3.1. Concentration of PTEs in the Ashes of Ryegrass Shoots

The concentrations of PTEs in ashes of ryegrass are summarized in Table 1. The average concentrations of PTEs in ashes ranged from 40 to 7543 mg kg<sup>−1</sup>. The concentrations of PTEs in the ryegrass were in the following ascending order: Pb < Cd < Cu < Fe < Al < Mn < Zn. It is worth mentioning that the concentrations of carcinogenic metals (Cd and Pb) were the lowest in comparison with those of metals, called “metals of interest” in the current study due to their application in a great number of organic reactions [35]. Among the studied PTEs, Zn appeared as the most phytoavailable. Since salts of Zn, Cu, Mn, Fe, and Al are potential Lewis acids (LA), the ratios of LA/Cd and LA/Pb values were calculated. As shown in Table 1, these ratios were 152.8 ± 5.5 and 188.3 ± 10.4, respectively. These results suggest that shoots of ryegrass may be considered as suitable biomass to synthesize molecules of interest via green chemistry and engineering processes after their conversion into biosourced catalysts (called Zn-BioCat).

**Table 1.** Concentration of potential toxic elements (PTEs) (mean ± standard deviation in mg kg<sup>−1</sup>; n = 6) in the ashes of ryegrass shoots and ratios [Zn]/[M].

| Element        | Cd          | Pb           | Zn         | Cu          | Mn        | Fe         | Al         |
|----------------|-------------|--------------|------------|-------------|-----------|------------|------------|
| Concentration  | 49 ± 1      | 40 ± 1       | 7543 ± 222 | 69 ± 4      | 908 ± 41  | 303 ± 8    | 597 ± 133  |
| Ratio [Zn]/[M] | 152.8 ± 5.5 | 188.3 ± 10.4 | 1          | 109.1 ± 9.4 | 8.3 ± 0.4 | 24.9 ± 1.0 | 13.2 ± 3.3 |

#### 3.2. Synthesis of Idrocilamide Using Zn-BioCat

As described in the previous section, idrocilamide was synthesized from methyl cinnamate (**1**), 2-aminoethanol (**2**), and Zn-BioCat. This reaction was conducted under a solvent-free condition at 70 °C for 24 h. Few studies reported the synthesis of this compound in the literature. Briefly, idrocilamide was first produced by Guan et al. [51] from 4 reactants using dichloromethane (CH<sub>2</sub>Cl<sub>2</sub>) as a solvent and then, by De Sarkar et al. [52] from 5 reactants in tetrahydrofuran (THF) as a solvent. Due to the high number of reactants, the use of solvents (CH<sub>2</sub>Cl<sub>2</sub> and THF), which are not classified as green solvents and their effects on human health [53,54], the proposed route was not compared to the reaction schemes of Guan et al. [51] and De Sarkar et al. [50]. In contrast, our method was compared to the procedure of Garg and Milton [55] in which the authors used the same reactants as those presented in Figure 1 and also sodium carbonate (Na<sub>2</sub>CO<sub>3</sub>) in methanol. The current strategy was also compared to the recent procedure described in Dufrénoy et al. [56] in which the authors used zinc chloride (ZnCl<sub>2</sub>) as a Lewis acid catalyst. These comparisons were made using metrics that are usually selected to evaluate the efficiency of chemical processes via the environmental and technical improvements. To mention only the most relevant metrics, yield, carbon efficiency, mass efficiency, environmental factor, and stoichiometric factor were calculated from the conditions used by the authors and those related to our procedure. The results are summarized in Table 2.

**Table 2.** Metrics associated with the determination of environmental factors.

| Method               | Yield (%) | Atom Economy (%) | Waste (%) | E-factor | RME (%) |
|----------------------|-----------|------------------|-----------|----------|---------|
| Garg and Milton [55] | 86.6      | 58.0             | 41.9      | 0.72     | 28.9    |
| Dufrénoy et al. [56] | 65.0      | 53.1             | 46.9      | 0.88     | 53.9    |
| This study           | 69.9      | 85.7             | 14.3      | 0.17     | 59.7    |

Yields of the reaction were between 65% and 86.6%, depending on the experimental conditions. Although the reaction yield from the current study was 5% higher than that obtained by Dufrénoy et al. [56], it was 16.7% lower than that from the study of Garg and Milton [52]. Nevertheless, it is worth mentioning that Zn-BioCat was reused after its liquid-solid filtration up to 5 runs without any significant loss of yield, which was up to 77%. This result fortifies the idea that it might be possible to use the unreacted reagents from one run to another or to conduct the reaction under continuous flow using a fixed bed plug-flow quartz reactor [57].

The atom economy of both routes described in the literature were 58.0% and 53.1% (Table 2). These results are mainly explained by the utilization of catalysts,  $\text{Na}_2\text{CO}_3$  [55] and  $\text{ZnCl}_2$  [56], both of them requiring a post-reaction treatment for their elimination. For instance,  $\text{ZnCl}_2$  treatment may present risks related to the release of HCl vapors resulting from the reaction between  $\text{ZnCl}_2$  and water during the hydrolysis process. In our procedure, the atom economy was 85.7%. This value is explained by the recycling and reuse of our Zn-BioCat without significant loss of activity and with reaction time-saving. The first consequence is the low production of waste using Zn-BioCat (14.3%) in comparison with the other two methods (41.9% and 46.9%, respectively). This may present a great economic interest for industries and may contribute to environment protection. The second consequence is related to the very low value of the E-factor, 0.17 for the proposed method (instead of 0.72 and 0.88 in the two others). This result confirms that our procedure fits into a context of waste minimization and resource efficiency in chemical manufacturing.

Finally, the reaction mass efficiency of Garg's reaction was 28.9%, which means that 71.1% of the weights of the reactants used in the reaction were not inserted in idrocilamide 3. In comparison with the study of Dufrénoy et al. [56], a 9.7% increase of RME was highlighted using Zn-BioCat instead of  $\text{ZnCl}_2$  (59.7% vs. 53.9%), indicating that 40.3% of reactants used in the reaction described in Figure 1 were not inserted in idrocilamide 3. With the same idea, the stoichiometric factor (SF) was equal to 1.75 for the reaction described by Garg and Milton [55], indicating that 75% of the weights of the reactants were used in excess by the authors. This metric was 1.03 and 1.00 in the process described by Dufrénoy et al. [56] and in the current study, respectively, demonstrating that no limiting and excess reactants were used in the process. It is worth noting that metrics described in Table 2 do not take into account the energy necessary for the thermal process related to the preparation of the ecocatalyst and to the synthesis of idrocilamide. In comparison with the methods described in the literature for the preparation of the target molecule [55,56], the energy costs are similar since the operatory conditions in terms of temperature and duration are analogous. In contrast, ash production needs high thermal energy demand. For this reason, contaminated ashes should be recovered from industries that used biomass to produce energy, thus enabling the heating of buildings. This should contribute to avoiding unnecessary use of energy for the preparation of ecocatalysts.

From a theoretical point of view, based on results obtained in the current study, 3 kg of ashes are necessary to produce 1 kg of idrocilamide per run. This corresponds to a mass of fresh shoot biomass of about 30 kg. In previous studies [58,59], heterogeneous catalysts from hyperaccumulators were used several times without any loss of catalytic activity. In the current study, Zn-BioCat was also reused 4 times, and it conserved its efficiency. Consequently, 5 kg idrocilamide should be produced by using 3 kg of contaminated ashes. Further studies are needed to inspect the recyclability of our Zn-BioCat in a scale-up production of idrocilamide (a semi-industrial process). Furthermore, it could be interesting to study the reactivity of the biosourced catalyst in the synthesis of idrocilamide derivatives.

#### 4. Conclusions

Shoots of ryegrass grown on contaminated garden soils by Cd, Pb, and Zn were transformed into a biosourced catalyst. This product is a Zn-enriched material since the average concentration of Zn in ashes is 7543 mg kg<sup>−1</sup>. The utilization of this catalyst in the synthesis of idroclamide was successful since the product was obtained in 69.9% yield under green conditions (solvent-free, without post-reaction treatment of the catalyst, reusable catalyst). Moreover, the limitation of waste, a maximization of material incorporation in the process, a minimization of the steps, and optimization of the stoichiometric factor were demonstrated by means of metrics related to the 12 principles of the green chemistry and engineering.

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