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Experimental investigation of Ethane and Propane injection under sub- and super-critical conditions

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Abstract

Studying a fluid flow under high-pressure conditions through reliable experiments is still nowadays a challenge. When the chamber pressure exceeds the critical pressure of working fluids the supercritical state of matter is reached and the distinction between gas and liquid becomes blurred. For such special conditions, experimental data are scarce and need to be consolidated. Indeed, the modification of the local refractive index induced by density gradient needs to be analyzed with appropriate image-based technique.

In the present study, the REFINE test bench (Real-gas Effect on Fluid Injection: a Numerical and Experimental study) has been designed at CORIA Lab to study the non-reactive injection of Ethane and Propane into Nitrogen under sub- and supercritical conditions. The ambient gas pressure can be raised up to 6 MPa and warmed up to 573 K to scan sub- and trans-critical injections. The chamber is equipped with two perpendicular optical axes allowing simultaneously different optical diagnostics. Experimental data are collected from shadowgraph and diffused backlight illumination techniques. Quantitative measurements of jet spreading angle and breakup length are compared to results coming from literature exhibiting multiple flows topologies.

Keywords

Supercritical, Experimental, Injection.

Introduction

In propulsion applications like in rocket engines, gas turbines or diesel engines when late injections are performed, injection often occurs at pressures that are above the critical point of the incoming fluids, i.e. at supercritical pressures [1, 2]. At such high-pressure conditions, injection is quite different from classical known behaviors observed at low pressure, and studying these new phenomena is challenging. Very few experiments dealing with non-reacting supercritical injection exists in the world, but a substantial experimental data set has been developed over the last few decades at DLR and AFRL laboratories [3]. The main issue is the ambient high pressure that locally affects the refractive index gradient making the delivery of experimental data highly challenging. Therefore this lack of reliable data is damaging for the development of models for turbulence and combustion [4, 5, 6].

The objectives of the research program REFINE (Real-gas Effects on Fluid Injection: a Numerical and Experimental study) are to bridge the gap between experimental and numerical observations in providing reliable experimental data and new strategies of modeling. The experimental aspect is presently addressed through the development of a new test-bench (next section) and for which a first qualitative analysis is presented. Indeed, many combustion devices involve thermodynamic phase transition from subcritical to supercritical state as in high-pressure combustion chambers of rocket or diesel engines. When the pressure exceeds the critical pressure p_c of any substance, the distinction between gas and liquid vanishes and the injection and combustion are quite different from classical known behaviors observed at low pressure. p_c (resp. T_c) is the highest pressure (resp. temperature) for which distinct liquid and gas phases can be observed [7]. Above these values spray atomization is changed [3, 8, 9]: drops are no longer observed on Shadowgraph or Schlieren images due to the reduced surface tension and heat of vaporization [10, 11]. Nevertheless a dense and dark core is still visible at the injector exit on shadowgraph [12]. The analysis of the spreading angle shows that supercritical jet has a behavior closer to a gaseous jet than to a liquid one [13, 14]. Only a few studies are dealing with experimental investigations of high-pressure injection and mixing process. Moreover, existing test benches often include additional effects that add a degree of complexity to a process already complex in nature. Indeed, particular conditions such as reactive flows, assisted-jet, cryogenic injection or acoustic interactions are often considered.

This paper introduces a new test bench and brings the first quantitative results that are compared to the literature. Injection of Propane was first studied but more investigations were explored for Ethane jets with in particular, the influence of the injection velocity, the temperature of the surrounding fluid, and the pressure. A first set of data on breakup length and spreading angle is also provided.

Experimental setup

The experiment has been designed to study injection and atomization of a non-assisted jet under high pressures conditions. REFINE is equipped with a steel chamber able to withstand pressures up to 10 MPa and moderate temperatures (573 K). Two optical axes through circular silica windows of $\varnothing 120$ mm are available for simultaneous measurements (see Fig. 1). The chosen working fluids are Ethane (C_2H_6 , $T_c = 305.3$ K and $p_c = 4.87$ MPa) and Propane (C_3H_8 , $T_c = 369.8$ K and $p_c = 4.25$ MPa). Their moderate critical pressure and temperature are within

the range of operating conditions achievable in the chamber. A one-liter pump supplies the working fluid at the liquid-gas equilibrium pressure. Before injection, the pump is filled up, and then the fluid contained inside the vessel of the pump is compressed to the target pressure plus 0.05 MPa to prevent the ambient fluid in the main chamber from entering the pump and creating an effervescent jet.

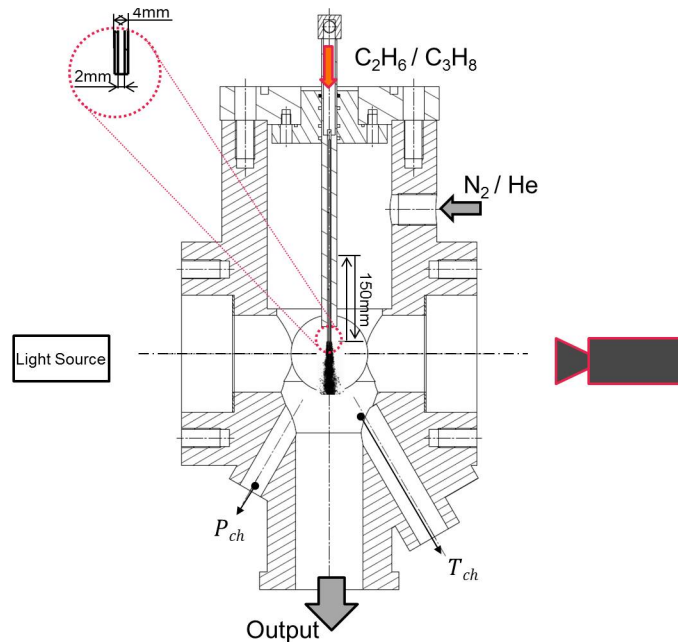


Figure 1. Section of the REFINE injection chamber.

The injection is performed with a sharp-edged stainless steel tube of 210 mm long, with a 4 mm outer diameter and a 2 mm inner diameter. The injector can be moved vertically over a 150 mm stroke by a motorized system located on the top of the chamber, to expand the field of view. The rig is fully instrumented with thermocouples, pressure gauges and mass flow meters as indicated in Fig. 2. The chamber is filled with Nitrogen and the pressure is controlled with a pressure regulator having a precision of ± 0.1 MPa. To warm up the ambient fluid, two MICA heating plates are placed at the bottom of the chamber and a heater (Fig. 2) is inserted just before the feeding of N_2 .

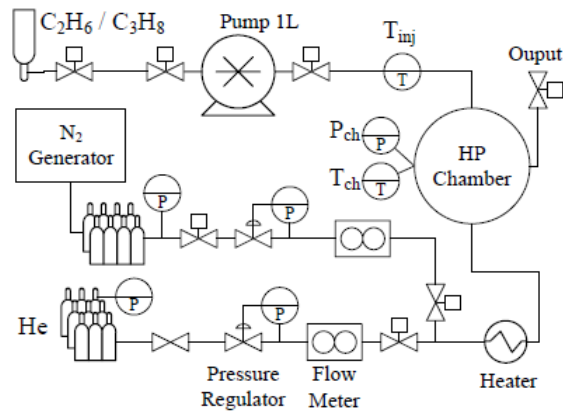


Figure 2. Flowchart of REFINE test bench.

Two optical diagnostics, diffused back-light illumination (DBI) and shadowgraphy, are used to quantify the injection and atomization of the working fluids. The main difference between these two techniques lies in the light source arrangement. For shadowgraphy, a collimated laser diode (Cavitar) is used whereas a stroboscope flashlamp (MVS-2601) with a glass diffuser is used for DBI. Depending on the situation and on the desired information, different camera configurations were used, as indicated in Tab. 1. The DBI allows detecting the largest density gradients, which are essentially located at the liquid-gas interface contrary to shadowgraphy that detects smoother density gradient.

All the cases studied hereafter are summarized in Tab. 2, which provides the values of pressure (P), temperature in the chamber (T_{ch}) or upstream the injector (T_{inj}) and velocity (V) calculated as the average of measurements made during the injection process. The \pm number indicates the variability of the measured quantity during this injection. The dimensionless numbers (Reynolds, Weber and Ohnesorge) are also provided.

Table 1. Table of the different configurations of camera and light source with the associate designation code used in this manuscript.

Code	Camera	Resolution (px/mm)	Field (mm)	Light
BG1	BlueCougar-x24	27.0	59.3 × 44.5	DBI
BG2-1	BlueCougar-x25a	53.5	45.8 × 38.1	DBI
BG2-2	BlueCougar-x25a	130.5	18.8 × 15.7	Shadowgraphy
PCO-1	PCO Edge 5.5	58.0	44.1 × 37.2	DBI
PCO-2	PCO Edge 5.5	121.5	294.3 × 17.8	DBI

Table 2. Tests case conditions with dimensionless numbers: Reynolds (Re), liquid and gaseous Weber (We_l , We_g) and Ohnesorge (Oh).

Figure	P (MPa)	V (m.s ⁻¹)	T _{ch} (K)	T _{inj} (K)	Re	We _l ¹	We _g ¹	Oh
Propane in Nitrogen at ambient temperature								
Fig. 3	2.01±0.05	2.12	295.3±0.4	294.3±0.9	20572	605	28	1.20E-3
		1.06			10286	151	7	
		0.53			5143	38	2	
		0.27			2620	10	0.45	
Ethane in Nitrogen at ambient temperature, indicated as T _{ch} = 293 K in the text								
Fig. 4,6,8,7,10	6.05±0.09	2.02	295.0±6.1	295.0±0.2	30545	3943	763	2.06E-3
	6.07±0.08	1.06	293.9±2.3	294.9±0.2	16073	1073	206	2.04E-3
	6.02±0.08	0.53	294.4±1.3	294.8±0.1	8048	269	51	2.04E-3
	6.03±0.09	0.27	295.1±1.2	294.8±0.1	4023	68	13	2.05E-3
	5.53±0.03	2.02	292.9±4.6	294.4±0.2	30878	3647	646	1.96E-3
	5.51±0.09	1.06	293.8±2.3	294.5±0.1	16175	1022	179	1.98E-3
	5.57±0.08	0.53	294.4±0.9	294.9±0.3	8152	267	48	2.01E-3
	5.66±0.01	0.27	295.1±0.2	295.4±0.1	4083	75	14	2.12E-3
	5.08±0.08	2.05	291.9±5.6	295.1±0.1	31966	3967	662	1.97E-3
	5.20±0.03	1.06	292.9±2.6	294.9±0.3	16513	1053	178	1.96E-3
	5.15±0.02	0.53	294.3±1.7	294.3±0.4	8215	245	41	1.91E-3
	5.12±0.01	0.27	294.2±1.1	295.4±0.1	4108	58	10	1.86E-3
	4.56±0.03	2.05	292.0±3.4	294.8±0.3	32491	3752	573	1.89E-3
	4.57±0.02	1.06	293.7±1.2	294.9±0.3	16890	1017	155	1.89E-3
	4.53±0.01	0.53	294.0±1.0	294.8±0.1	8450	251	38	1.88E-3
	4.50±0.00	0.27	293.9±0.4	295.2±0.1	4250	65	10	1.90E-3
	4.09±0.07	2.05	291.5±4.7	294.7±0.3	33185	3610	509	1.81E-3
	4.07±0.04	1.06	293.4±1.3	294.5±0.3	17212	953	132	1.79E-3
	4.07±0.04	0.53	294.1±0.5	294.3±0.2	8583	235	32	1.79E-3
	4.08±0.02	0.27	294.0±1.3	294.3±0.1	4294	58	8	1.78E-3
Fig. 5	6.02±0.01	0.53	294.4±1.1	294.0±0.3	7988	245	46	1.96E-3
	6.01±0.01	0.27	295.0±0.7	294.0±0.3	3995	61	11	1.96E-3
	5.53±0.05	0.53	294.5±0.8	293.9±0.5	8090	239	42	1.91E-3
	5.50±0.04	0.27	294.9±0.5	293.9±0.5	4049	60	10	1.91E-3
	5.05±0.08	0.53	294.4±1.0	293.7±0.5	8196	231	37	1.85E-3
	5.06±0.08	0.27	294.7±0.5	293.8±0.4	4100	58	9	1.86E-3
	4.56±0.07	0.53	294.4±0.9	293.5±0.5	8322	222	33	1.79E-3
	4.55±0.07	0.27	294.9±0.8	293.6±0.5	4167	56	8	1.80E-3
	4.08±0.06	0.27	294.7±0.4	293.1±0.6	8490	215	29	1.73E-3
	4.06±0.06	0.53	294.4±0.9	293.4±0.3	4226	52	7	1.71E-3
Ethane in Nitrogen at T _{ch} > T _c , indicated as T _{ch} = 323 K in the text								
Fig. 8,7,10	6.05±0.03	2.02	326.4±6.9	298.4±0.2	31480	6309	1108	2.52E-3
	6.06±0.01	1.06	328.0±0.9	298.2±0.1	16535	1674	292	2.47E-3
	6.06±0.01	0.27	328.9±0.3	298.1±0.1	8260	418	73	2.47E-3
	6.06±0.01	0.53	329.5±0.1	298.4±0.1	4141	110	19	2.53E-3
	5.58±0.03	2.02	327.9±6.1	297.9±0.0	31831	5656	924	2.36E-3
	5.51±0.00	1.06	329.0±3.3	298.2±0.2	16843	1648	267	2.41E-3
	5.54±0.01	0.27	329.7±0.7	297.4±0.1	8349	366	59	2.29E-3
	5.51±0.01	0.53	329.5±0.3	298.1±0.1	4207	102	16	2.40E-3
	5.11±0.02	2.04	326.5±6.9	297.9±0.1	32834	5682	873	2.30E-3
	5.08±0.02	1.06	332.7±4.5	300.0±0.3	17483	2268	352	2.72E-3
	5.08±0.00	0.27	332.1±1.1	298.4±0.1	8585	416	63	2.38E-3
	5.09±0.00	0.53	332.1±0.1	298.1±0.1	4277	100	15	2.34E-3

¹ Considering pure Ethane at T_{inj} position and taken a surface tension from NIST (<http://webbook.nist.gov/chemistry/>) in liquid phase on the evaporation line at the T_{inj}

Results and discussion

The behavior of a cylindrical jet discharging into a quiescent atmosphere is commonly characterized [15] by the jet stability curve as shown in Fig. 3. This curve shows the variation of the jet breakup length (L_{bu}) as a function of the discharge velocity. The breakup length is defined as the distance between the injection orifice and the first discontinuity in the liquid jet. Five flow regimes are identified. They are respectively named the dripping regime ($a \rightarrow b$), Rayleigh ($b \rightarrow c$), first wind-induced ($c \rightarrow d$), second wind-induced ($d \rightarrow e$) and atomization regimes ($> e$). The dripping regime is not of interest here.

Images of a Propane jet discharging into an atmosphere of nitrogen at 293 K and 2 MPa are presented along the stability curve in Fig. 3. Refraction of light near the liquid-gas interface yields to steep grey level gradient making easier the localization of the interface. Brighter areas in the center are also due to the refraction of light which is less deviated when going through the jet axis. The more the surface is corrugated, the more deviated is the light making the jet darker. For $V = 0.25 \text{ m.s}^{-1}$, the jet morphology is characteristic of the Rayleigh regime. Increasing the velocity up to 0.5 m.s^{-1} , the breakup length becomes greater than the field of view. The discharge velocity then approaches the critical velocity corresponding to point c . For $V = 1 \text{ m.s}^{-1}$ the breakup length becomes shorter, confirming that the critical velocity has been exceeded. The surface of the jet becomes corrugated and the first wind-induced regimes is reached. For $V = 2 \text{ m.s}^{-1}$, disturbances on the jet surface are getting stronger and L_{bu} is no longer observable within this field of view, i.e. the point d is passed and the jet is in the second wind-induced regime.

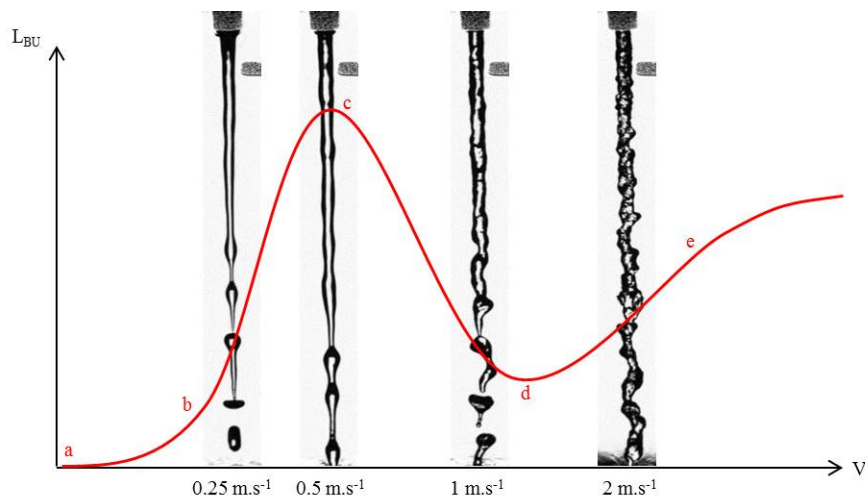


Figure 3. Injection of Propane at 2.0 MPa and ambient temperature along the stability curve [15](BG1).

Images of a jet of Ethane into Nitrogen at 293 K and 6.0 MPa are shown in Fig. 4. The breakup length of Ethane jet has been investigated for $V = 0.25 \text{ m.s}^{-1}$ and $V = 0.5 \text{ m.s}^{-1}$. For $V \geq 1 \text{ m.s}^{-1}$, the atomization starts only a few diameters downstream the injection orifice and the cloud of drops around the jet makes difficult the estimation of L_{bu} . The intact core length is thus commonly used in these conditions to characterize the jet atomization [16].

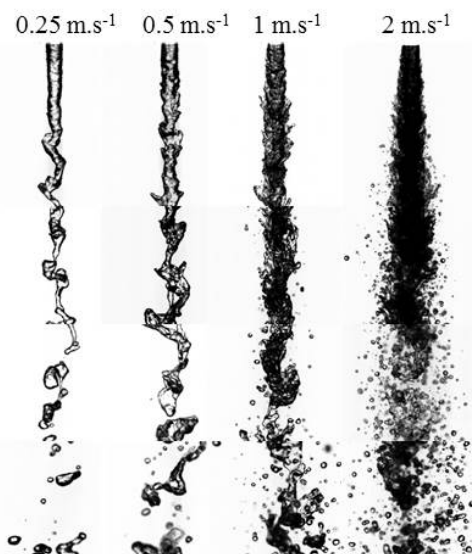


Figure 4. Injection of Ethane at 6.0 MPa and ambient temperature (PCO-2).

Statistics on L_{bu} of the Ethane jet was estimated for the two lowest discharge velocities. Since the jet deviates from its central axis, both perpendicular visual axes were used (BG2-1 & PCO-1) to make sure that no drop or fluid deformations hide the first discontinuity of liquid. 600 two-level images were taken and analyzed for each case to compute the probability density functions (pdf) for L_{bu} (see Fig. 5). For all the studied ambient pressures, the slowest velocity gives the shortest breakup length, meaning that L_{bu} increases with the discharge velocity. Since the shape of the jet (see Fig. 4) has nothing in common with a Rayleigh jet, those jets can be classified in the second wind-induced regime. It can also be observed that the pdf of L_{bu} are wider for a larger velocity and for $V = 0.25 \text{ m.s}^{-1}$ the maximum peak is better defined. No conclusion could be drawn on the influence of pressure on the breakup length.

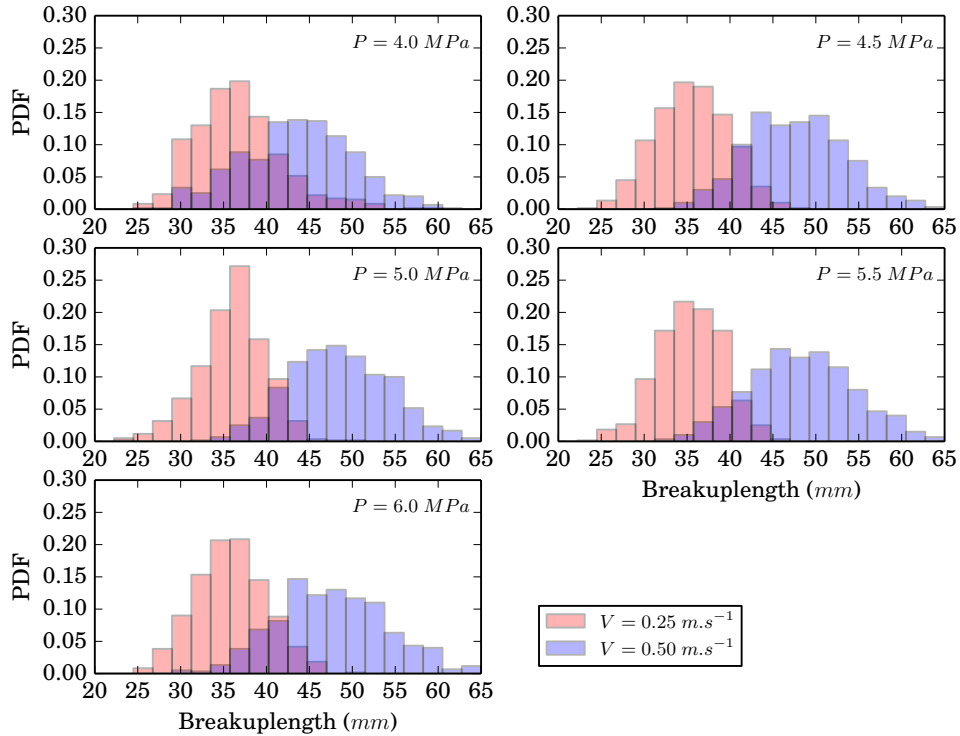


Figure 5. PDF of 600 breakup lengths of Ethane jet at 293 K. Pressure from 4.0 MPa to 6.0 MPa.

The impact of the chamber pressure on jet disintegration is illustrated in Fig. 6 for Ethane injected into N_2 at 293 K with a fixed discharge velocity of 1 m.s^{-1} . The pressure ranges from 4.0 MPa to 6.0 MPa, i.e. from subcritical to supercritical pressure if referring to Ethane critical pressure. Increasing the pressure leads to more perturbed jets with smaller drops, similarly to the observations done by Oswald [17] for an injection of N_2 at 5 m.s^{-1} into N_2 at rest.

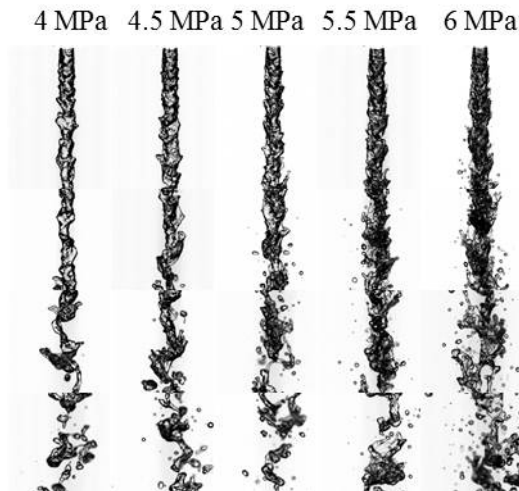


Figure 6. Injection of Ethane into Nitrogen at $V = 1 \text{ m.s}^{-1}$ for different levels of ambient pressure with $T = 293 \text{ K}$. 4 images are used to create a field of visualization of 66.1 mm length (PCO-2).

Table 3. Measured temperatures T_{ch} , T_{inj} and T_{jet} in the case of an injection of Ethane into N_2 for different velocities.

P (MPa)	V ($m.s^{-1}$)	T_{ch} (K)	T_{inj} (K)	T_{jet} (K) at 6 mm	State ²
6.0	0.25	327	298	308	Supercritical
6.0	0.50	325	297	303	Liquid
6.0	1.00	323	298	300	Liquid
6.0	2.00	328	298	298	Liquid

To explore the effect of the ambient temperature, various configurations of Ethane jets into Nitrogen for four discharge velocities and a temperature of N_2 set to $T = 293\text{ K} < T_c$ or $T = 323\text{ K} > T_c$ are shown in Fig. 7. Pressure is fixed to 6.0 MPa, i.e. above the critical pressure of Ethane. The temperature of the jet (T_{jet}) measured in the fluid at the nozzle exit with an exposed junction thermocouple placed 6 mm downstream from the nozzle exit are summarized in Tab. 3. These temperature measurements were done on different experiments, i.e. not simultaneously to image acquisition. This temperature, T_{jet} , decreases with the increase of the velocity. The discharge velocity must be considered as a key-point parameter since it induces modification in the ratio between characteristic convection and thermal times

For the lowest injection velocities ($V = 0.25\text{ m.s}^{-1}$ and $V = 0.5\text{ m.s}^{-1}$), the temperature has a clear impact on jet disintegration since T_{jet} is close to the critical temperature of Ethane ($T_c = 305\text{ K}$). For example, a diffuse jet with dense pockets appears for $T_{ch} = 323\text{ K}$ and $V \leq 0.5\text{ m.s}^{-1}$ whereas a liquid-gas interface is clearly visible for $T_{ch} = 293\text{ K}$.

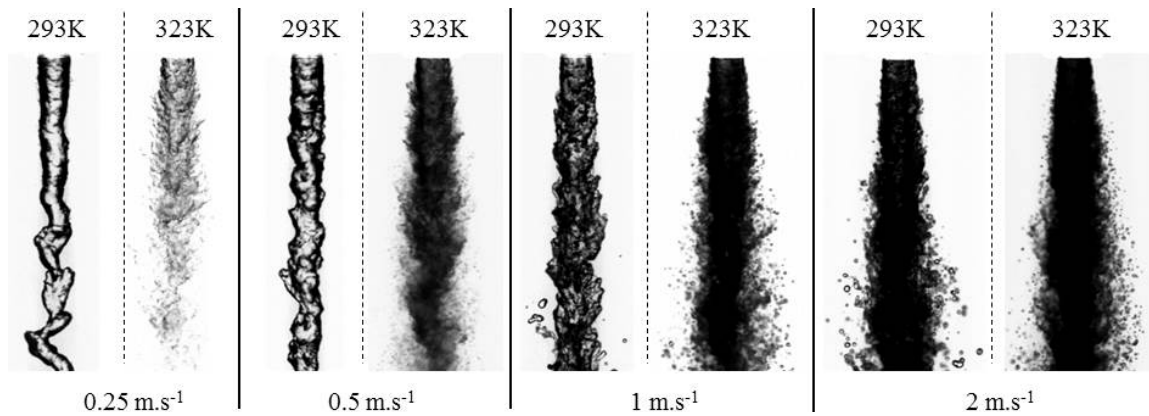


Figure 7. Jets of Ethane in Nitrogen at 6 MPa (PCO2).

Increasing the velocity to $V = 1\text{ m.s}^{-1}$, small fluid elements identified as drops appear whatever T_{ch} . However, there are more droplets identified for $T = 323\text{ K}$. The same observation is done for $V = 2\text{ m.s}^{-1}$. Hence, having a temperature of the ambient fluid above the critical temperature of the injected fluid strongly modifies the jet behavior. The interaction between the injected fluid and the ambient one is greater for low injection velocities where a diffusion process is observed, but modifies also the atomization process at higher injection velocity as stated in [3, 18].

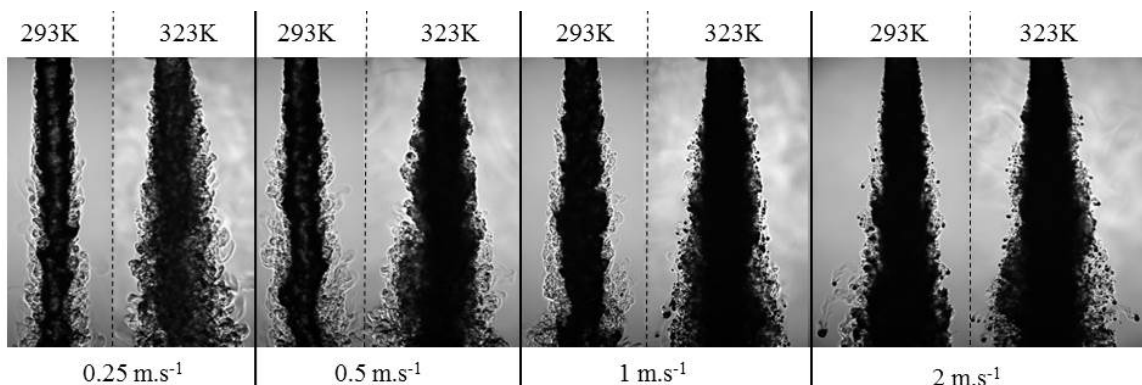


Figure 8. Jets of Ethane in Nitrogen at 6 MPa (BG2-2).

With DBI visualizations some finger-like structures are observed at the jet interface. In order to analyze those structures, shadowgraph visualizations were achieved to explore the density gradient around the liquid jet (or around the dense core). Accordingly, shadowgraph images (Fig. 8) have been taken simultaneously to DBI images (Fig. 7) on the perpendicular axis. A grey level gradient clearly appears around the liquid. It consists of a $C_2H_6-N_2$

²Considering pure Ethane at the T_{jet} position.

mixing layer varying radially from pure Ethane in the center of the jet to pure Nitrogen far from the jet. Two topologic parameters are then defined from shadowgraph images: the dark-core angle and the spreading angle also named growth rate [19]. From the normalized images of raw shadowgraphs, a two-level image is created. The black and white delimitation gives an interface between dense and light fluid. The calculation of the standard deviation of 120 two-level shadowgraph images yields the probability of presence of this interface on a grey-level scale. Only the values above 90 % of the maximum of this probability are retained to define the area delimited by the dark-core angle θ_d (inside boundary in red on Fig. 9) and the spreading angle θ_s (outside boundary in green). The spreading angle for different pressures, velocities and chamber temperatures of an Ethane jet are plotted in Fig. 10.

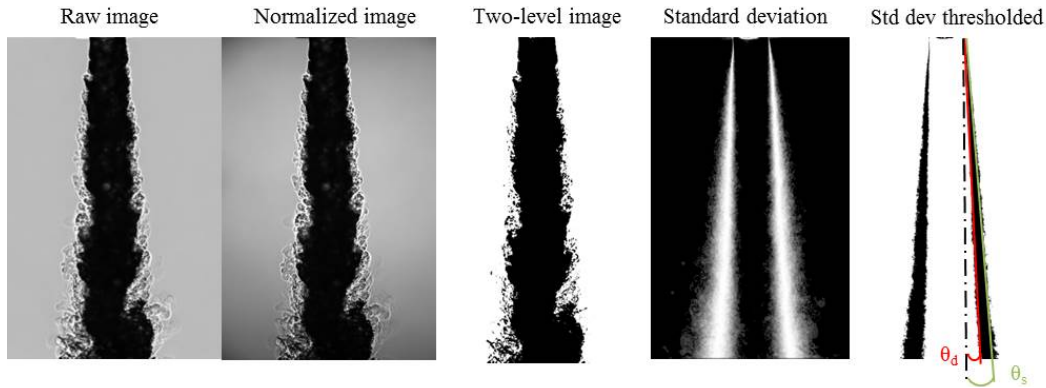


Figure 9. Image processing from raw image to the threshold of the standard deviation. On the last image: the dark-core angle θ_d is on the inside boundary in red and the spreading angle θ_s is outside in yellow.

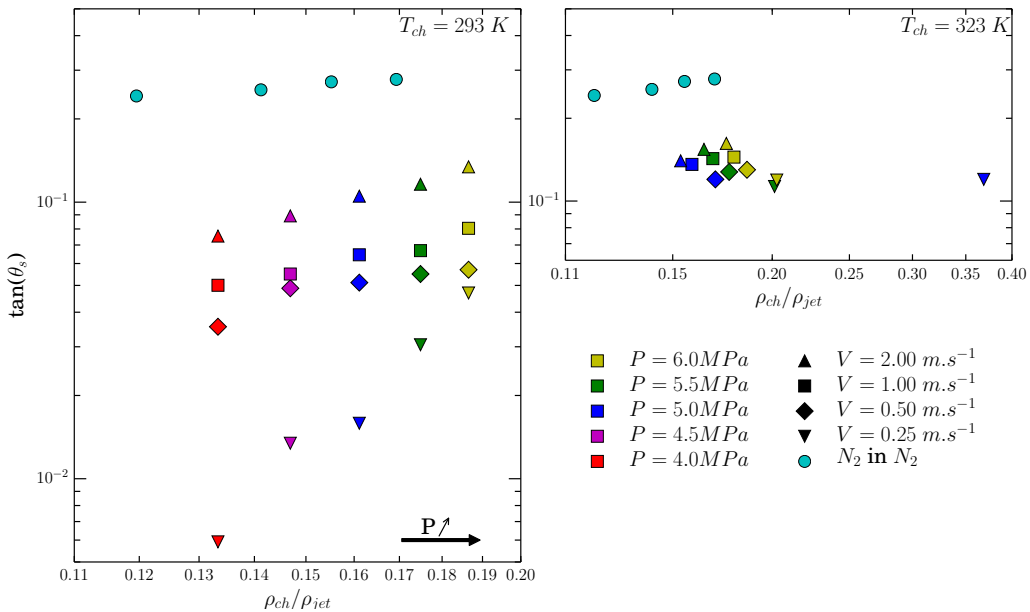


Figure 10. Spreading angle versus the chamber to injectant density ratio. For comparison, spreading angle of liquid N_2 injection in N_2 at ambient temperature from Fig. 8 of Chehroudi's paper [12].

To be compared to the results summarized by Chehroudi in [12], the spreading angle is plotted against the ambient over injectant density ratio. The density of pure fluids is based on the measured temperature (Tab. 3) and estimated from NIST tables. For example, if $T_{ch} = 323 \text{ K}$, then the temperature of Ethane considered for NIST tables is T_{jet} . For each velocity, the growth rate increases with the pressure (also with the density ratio), which is in agreement with the other studies. At a given ambient pressure, the spreading angle increases with the velocity. Ethane spreading angles are smaller to the AFRL data for liquid N_2 in N_2 [12] having similar density ratios. This difference can be due to the velocity range in our experiments, which is one of the lowest in the literature but no indication of velocity is given on the original plot of Chehroudi (Fig. 8 in [12]). Increasing the temperature makes a significant change of flow topology at low velocities. In Fig. 10 (right), the injection of Ethane at $V = 0.25 \text{ m.s}^{-1}$ under a pressure of 5.0 MPa exhibits a density ratio of $\rho_{ch}/\rho_{jet} = 0.35$ contrary to configurations at higher pressure ($\rho_{ch}/\rho_{jet} \approx 0.20$). Indeed decreasing the pressure makes the working conditions closer to the critical point. The limit of plotting the spreading angle against the density ratio is that when the chamber is heated, the jet temperature varies along the jet. There is even a temperature gradient in Nitrogen evidenced by the non-uniform grey level in shadowgraph images far from the jet at 323 K in Fig. 8.

Conclusions

A new fully instrumented experimental test-bench has been developed at CORIA lab to study the injection of inert fluids under supercritical conditions. In this study, Ethane and Propane are injected into a cold or warm environment of Nitrogen. The knowledge of the behavior of those fluids has been detailed. Results for the breakup length and the spreading angle are in agreement with the literature. It has been shown that the temperature of the ambient fluid played a major role in jet disintegration when pressure exceeds the critical pressure of the injected fluid. In case of an ambient temperature higher than the critical temperature of injected fluids a diffuse mixing process occurs for low discharge velocities. For a higher injection velocities, a process of atomization appears but with a larger spatial distribution of drops. These results were obtained by shadowgraphy and diffused backlight illumination that gave complementary information. Other optical diagnostics as schlieren and CBOS will be soon implemented to complement these first results. A X-ray analysis is currently under progress to deliver a density gradient measurement and to give additional characterization of the jet behaviors.

Acknowledgements

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